

5.1 The nuclear model of the atom

Name: _____ Class: _____ Date: _____

Total: 10 marks

KEY WORDS neutron 中子 • proton 质子 • electron 电子 • nucleus 原子核 • isotope 同位素

Objective

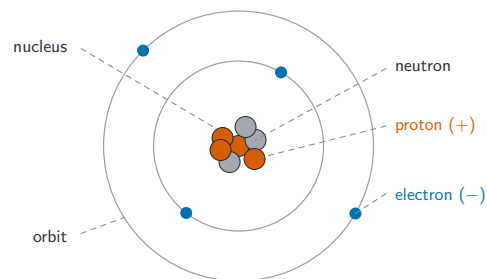
Build the skills to answer exam questions on the **nuclear model of the atom** 原子核模型 — atomic structure, the **alpha-scattering** α 散射 evidence, **nuclide notation** 核素符号, and **isotopes** 同位素.

You must be able to:

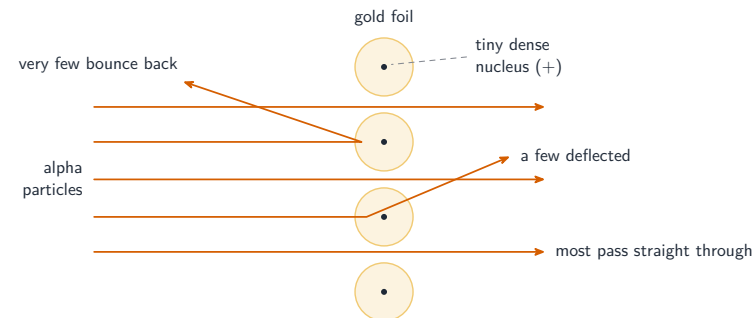
- describe the nuclear atom (tiny dense nucleus, electrons in orbits)
- interpret the alpha-scattering results
- use ${}_Z^AX$ to find protons, neutrons and electrons; define isotopes

1 Worked examples

■ Nuclide notation & scattering



Almost all the mass is in the tiny nucleus; the atom is mostly empty space



Most pass straight through (empty space); a few bounce back (small, heavy, positive nucleus)

${}_{79}^{197}\text{Au}$: 79 protons, $197 - 79 = 118$ neutrons, 79 electrons (neutral atom).

2 Practice

2.1 State what is meant by an isotope. [1]

2.2 For ${}_{11}^{23}\text{Na}$, write down the number of protons and the number of neutrons. [2]

3 Exam-style questions

3.1 In the alpha-scattering experiment, most alpha particles pass straight through the foil. This shows the atom is: [1]

A	B
C	D

- A solid throughout
- B mostly empty space
- C negatively charged
- D very large

3.2 The nuclide $^{40}_{19}\text{K}$ is given.

(a) State the number of protons, neutrons and electrons in a neutral atom. [3]

(b) An isotope of potassium is $^{39}_{19}\text{K}$. State how its nucleus differs from $^{40}_{19}\text{K}$. [1]

3.3 Explain why a very few alpha particles bounce almost straight back off the gold foil.[2]
