IGCSE Chemistry: **w25 22**

Multiple Choice Answers

Q17: B

IGCSE Chemistry: **w25 23**

Multiple Choice Answers

Q16: C

IGCSE Chemistry: **s25 22**

Multiple Choice Answers

Q16: D

IGCSE Chemistry: **s25 23**

Multiple Choice Answers

Q17: A

IGCSE Chemistry: **s25 42**

1(a)	78(%)
1(b)	450
1(c)	3
1(d)	2
1(e)	any value between 25 to 35 inclusive
1(f)	2
1(g)	8
1(h)	7

IGCSE Chemistry: **m25 42**

1(a)	21
1(b)	450
1(c)	2
1(d)	8
1(e)	16
1(f)	6.02×10^{23}

4(a)	no reactants or products can enter or leave
4(b)	enthalpy change (of reaction)
4(c)	(The value of ΔH is) negative
4(d)	**+**105 (kJ / mol)
4(e)	decreases increases increases no change
4(f)	**M1** Bond energy in breaking bonds **=** 1075 + 240 = **(+) 1315** **M2** Bond energy of $COCl_2$ = 1315 – (–105) = 1420 **M3** C=O = 1420 – 2 × C–C*l* = 1420 – (2 × 340) = (+) 740
4(g)	**M1** both single bonding dot and cross pairs correct **M2** double C=O bond dot and cross pairs are correct **M3** Complete diagram is correct