

2(a)	strontium / Sr
2(b)	M1 metallic M2 positive ions / cations M3 sea of electrons / mobile electrons / delocalised electrons M4 attraction between positive ions and electrons
2(c)(i)	M1 effervescence M2 solid dissolves M3 solution turns blue
2(c)(ii)	calcium hydroxide (1) hydrogen (1)
2(d)(i)	orange-red
2(d)(ii)	$2\text{Ca} + \text{O}_2 \rightarrow 2\text{CaO}$ M1 CaO as only product M2 correct equation
2(e)(i)	water(s) of crystallisation
2(e)(ii)	M1 $3.28 \div 164 = 0.02(00)$ M2 $1.44 \div 18 = 0.08(00)$ M3 $= \text{M2} \div \text{M1} = 4$

5(a)(i)	$\text{Pb}^{2+}(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) \rightarrow \text{PbSO}_4(\text{s})$ M1 PbSO ₄ as the ONLY formula on the right M2 ONLY Pb ²⁺ and SO ₄ ²⁻ on the left M3 state symbols
5(a)(ii)	lead(II) nitrate
5(a)(iii)	sodium sulfate / potassium sulfate / ammonium sulfate
5(a)(iv)	residue
5(a)(v)	M1 wash with distilled water M2 suitable method of drying
5(b)(i)	<i>any two from:</i> - good conductor of electricity - soluble salts - basic oxides
5(b)(ii)	<i>any two from:</i> - variable oxidation states - coloured compounds - catalyst