IGCSE Chemistry: **w25 21**

Multiple Choice Answers

Q10: D

Q11: C

IGCSE Chemistry: **w25 22**

Multiple Choice Answers

Q10: D

IGCSE Chemistry: **w25 23**

Multiple Choice Answers

Q12: C

IGCSE Chemistry: **w25 41**

3(a)	**M1** $2H_2O$ **M2** state symbols (aq) and (l)
3(b)	neutralisation
3(c)	$0.0100 \times 1000 \div 40 = 0.25(0)$
3(d)	burette
3(e)	**M1** mol H_2SO_4 needed = $0.0100 \div 2 = 0.005(00)$ **M2** vol of H_2SO_4 needed = $0.005(00) \div 0.200 \times 1000 = 25$ (cm^3)
3(f)	**M1** yellow **M2** orange
3(g)	repeat steps 1 and 3 without indicator
3(h)(i)	**M1** (a solution containing the) maximum concentration of a solute dissolved in the solvent **M2** at a specified temperature
3(h)(ii)	solubility decreases as temperature decreases
3(h)(iii)	none

IGCSE Chemistry: **w25 42**

3(a)	→ $BaSO_4(s)$ + 2NaCl(aq) **M1** 2NaCl (only) **M2** (s) and (aq)
3(b)	**M1** M_r $BaCl_2 = 208$ **M2** $208 \times 0.1 = 20.8$
3(c)	$0.100 \div 1.25$ ($\times 1000$) = 80 (cm^3)
3(d)	white
3(e)	residue
3(f)	**M1** rinsing the residue (with water) **M2** (mass was greater) because (dry) residue contains solid sodium chloride / NaCl adds to the mass (of residue)
3(g)	precipitation
3(h)	**M1** named soluble barium **salt** other than barium chloride e.g. (barium) nitrate **M2** named soluble sulfate **salt** other than sodium sulphate e.g. potassium (sulfate) / ammonium (sulfate) **M3** named insoluble barium **salt** other than barium sulfate e.g. (barium) carbonate

IGCSE Chemistry: **s25 21**

Multiple Choice Answers

Q9: A

Q20: B

IGCSE Chemistry: **s25 22**

Multiple Choice Answers

Q9: A

IGCSE Chemistry: **m25 22**

Multiple Choice Answers

Q12: C