

- 6** A sample of copper has a relative atomic mass of 63.5.

What are the relative abundances of the isotopes in this sample of copper?

- A** 25% ^{63}Cu and 75% ^{65}Cu
- B** 50% ^{63}Cu and 50% ^{65}Cu
- C** 75% ^{63}Cu and 25% ^{65}Cu
- D** 90% ^{63}Cu and 10% ^{65}Cu

- 5** Element T has two isotopes, $^{12}_6\text{T}$ and $^{14}_6\text{T}$.

Which statement about these isotopes is correct?

- A** They have different chemical properties because they have different numbers of neutrons.
- B** They have the same chemical properties because they have the same number of outer shell electrons.
- C** They have the same nucleon number because the sum of the number of protons and electrons is the same.
- D** They have different positions in the Periodic Table because they have different numbers of neutrons.