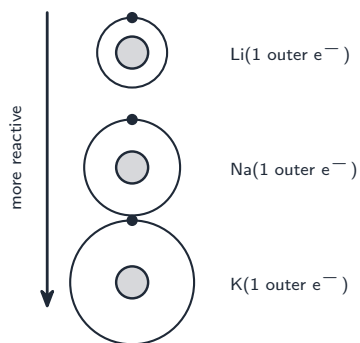




Watch out: elements in a group share properties because they have the **same number of outer electrons** (the group number), not the same total. Group I gets **more** reactive *down* the group — the opposite of Group VII.

Practice

Try these in order.

8.1 In what order are the elements arranged in the Periodic Table? [1]

8.2 What are (a) the rows, (b) the columns called? [2]

8.3 Why do elements in the same group have similar chemical properties? [2]

8.4 What charge of ion does a Group I element form? [1]

8.5 State the trend in reactivity going down Group I. [1]

8.6 Challenge. Potassium is below sodium in Group I. Predict whether potassium is more or less reactive than sodium, and explain why in terms of the outer electron. [3]