

6.2 Rate of reaction

Name: _____ Class: _____ Date: _____

Total: 10 marks

KEY WORDS concentration 浓度 • rate of reaction 反应速率 • activation energy 活化能
• collision theory 碰撞理论 • particles 粒子

Objective

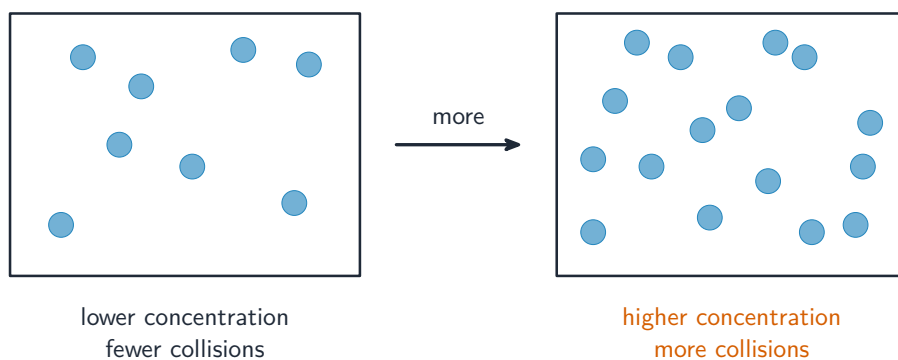
Build the skills to answer exam questions on the **rate of reaction** 反应速率 — **collision theory** 碰撞理论, the factors that change the rate, **catalysts** 催化剂, and measuring rate.

You must be able to:

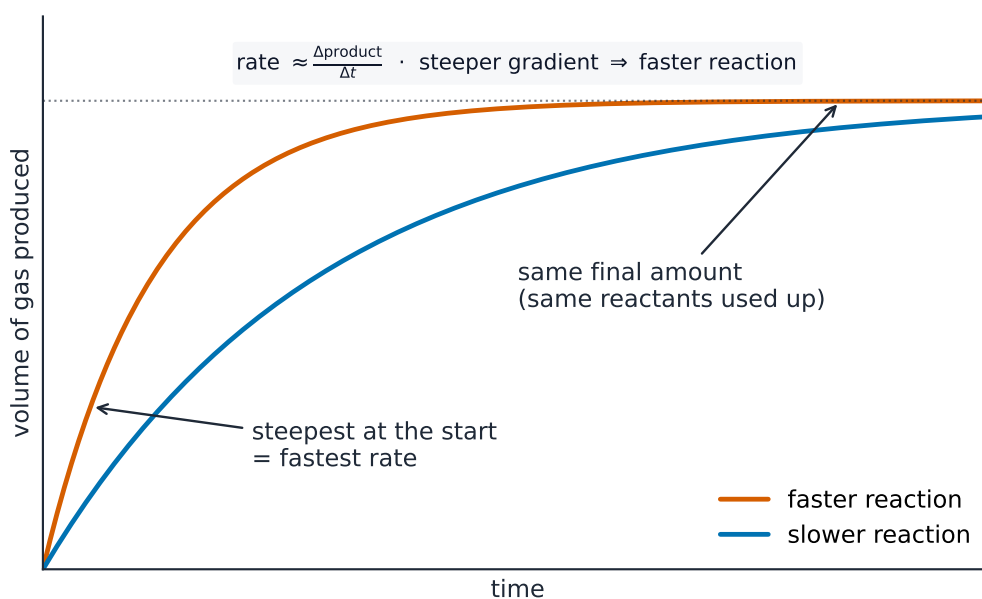
- use collision theory to explain the effect of concentration, surface area, temperature, pressure
- explain how a catalyst speeds up a reaction
- interpret a rate graph (steepest at the start, levels off)

1 Worked examples

■ Collision theory & rate graphs



More particles in the same volume collide more often → faster rate



Rate is fastest at the start (steepest), then levels off when the reaction finishes

- **catalyst:** lowers the activation energy → more successful collisions; not used up.

2 Practice

2.1 State two factors that increase the rate of a reaction.

[2]

2.2 State how a catalyst speeds up a reaction.

[1]

3 Exam-style questions

3.1 Increasing the temperature speeds up a reaction mainly because the particles:

[1]

A	B
C	D

A become larger

B move faster and collide more often with enough energy

C become catalysts

D stop colliding

3.2 Marble chips react with hydrochloric acid, giving off carbon dioxide.

(a) Explain, using collision theory, why powdered marble reacts faster than large chips. [2]

(b) Describe one method to measure the rate of this reaction. [2]

3.3 On a graph of gas volume against time, explain why the line becomes less steep and then flat. [2]
