

5.1 Exothermic and endothermic reactions

Name: _____ Class: _____ Date: _____ Total: 8 marks

KEY WORDS reaction pathway diagram 反应进程图 • reactants 反应物 • bond energy 键能
• enthalpy change 焓变 • products 生成物

Objective

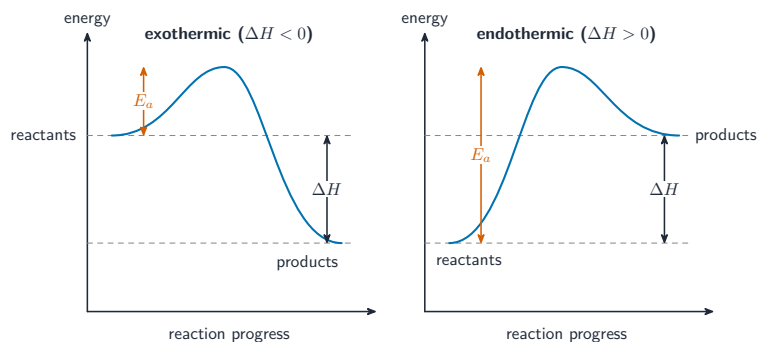
Build the skills to answer exam questions on **exothermic and endothermic reactions** 放热与吸热反应 — temperature changes, **enthalpy change** 焓变 ΔH , **reaction pathway diagrams** 反应进程图, and **bond energy** 键能 calculations.

You must be able to:

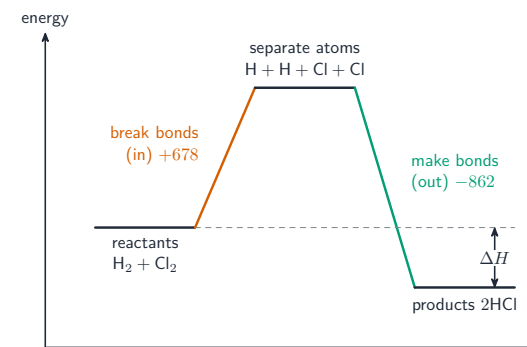
- classify reactions as exothermic ($\Delta H < 0$) or endothermic ($\Delta H > 0$)
- interpret a reaction pathway diagram (activation energy, ΔH)
- calculate ΔH from bond energies

1 Worked examples

■ Bond energy calculation



Exothermic: products lower ($\Delta H < 0$). Endothermic: products higher ($\Delta H > 0$)



$$\text{net: } \Delta H = 678 - 862 = -184 \text{ kJ/mol (exothermic)}$$

$$\Delta H = (\text{bonds broken}) - (\text{bonds made})$$

$\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$: broken $436 + 242 = 678$; made $2 \times 431 = 862$; $\Delta H = 678 - 862 = -184 \text{ kJ/mol}$.

2 Practice

2.1 State the sign of ΔH for an exothermic reaction. [1]

2.2 State whether bond breaking takes in or gives out energy. [1]

3 Exam-style questions

3.1 During an endothermic reaction, the temperature of the surroundings: [1]

- | | |
|---|---|
| A | B |
| C | D |
- A rises
B falls
C stays the same
D rises then falls

3.2 Consider $\text{H}_2 + \text{F}_2 \rightarrow 2\text{HF}$. Bond energies (kJ/mol): $\text{H-H} = 436$, $\text{F-F} = 158$, $\text{H-F} = 568$.

(a) Calculate ΔH . [3]

(b) State whether the reaction is exothermic or endothermic. [1]

3.3 On a reaction pathway diagram, state what the height of the 'hill' above the reactants represents. [1]
