

4.1 Electrolysis

Name: _____ Class: _____ Date: _____

Total: 13 marks

KEY WORDS electrodes 电极 • anode 阳极 • aqueous 水溶液 • electrolysis 电解 • cathode 阴极

Objective

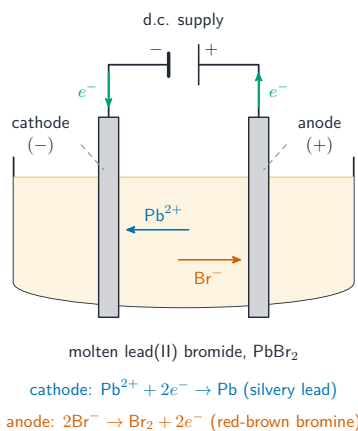
Build the skills to answer exam questions on **electrolysis** 电解 — the **electrolytic cell** 电解池, products at each electrode, **half-equations** 半反应式, and **electroplating** 电镀.

You must be able to:

- name the anode/cathode and explain why the substance must be molten/aqueous
- predict the products and write half-equations (reduction at cathode, oxidation at anode)
- describe electroplating

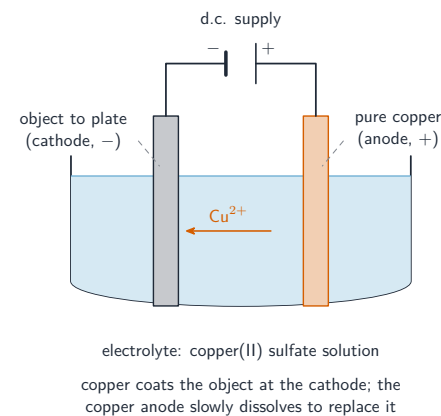
1 Worked examples

■ Electrolysis of molten PbBr_2



Positive ions go to the cathode (reduction), negative ions to the anode (oxidation)

- cathode (–):** $\text{Pb}^{2+} + 2\text{e}^- \rightarrow \text{Pb}$ (reduction).
- anode (+):** $2\text{Br}^- \rightarrow \text{Br}_2 + 2\text{e}^-$ (oxidation).
- needs molten/aqueous so the ions are free to move.



To electroplate: object = cathode, plating metal = anode, in a solution of the plating-metal ions

2 Practice

2.1 State why solid sodium chloride cannot be electrolysed. [1]

2.2 Name the products at each electrode when molten lead(II) bromide is electrolysed. [2]

3 Exam-style questions

3.1 At the cathode during electrolysis, positive ions: [1]

A	B
C	D

- A lose electrons (oxidation)
- B gain electrons (reduction)
- C stay unchanged
- D become anions

3.2 Dilute sulfuric acid is electrolysed using inert carbon electrodes.

(a) Name the gas formed at each electrode, and state the ratio of their volumes. [3]

(b) Write the half-equation for the reaction at the cathode, and explain why the ions in the water are discharged rather than the sulfate ions. [3]

3.3 Describe how you would electroplate an iron spoon with silver. [3]
