

12.5 Identification of ions and gases

Name: _____ Class: _____ Date: _____

Total: 21 marks

KEY WORDS sodium 钠 • nitrate 硝酸盐 • barium 钡 • sodium hydroxide 氢氧化钠
• nitric acid 硝酸

Objective

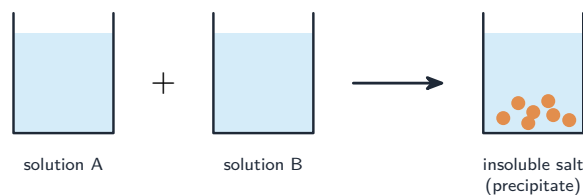
Build the skills to answer exam questions on the **identification of ions and gases** 离子与气体检验 — tests for **anions** 阴离子, **cations** 阳离子, gases, and **flame tests** 焰色试验.

You must be able to:

- give the test and result for common anions (carbonate, halides, sulfate, nitrate)
- give the cation tests with sodium hydroxide / aqueous ammonia
- give the gas tests and flame-test colours

1 Worked examples

■ Key tests



Many ion tests give a coloured precipitate

- **chloride:** dilute HNO_3 + silver nitrate $\rightarrow$ white precipitate.
- **sulfate:** dilute HNO_3 + barium nitrate $\rightarrow$ white precipitate.
- **CO_2 :** turns limewater milky. **H_2 :** squeaky pop. **O_2 :** relights a glowing splint.
- **flame:** Li red, Na yellow, K lilac, Cu blue-green.

2 Practice

2.1 Give the test and result for carbon dioxide gas. [2]

2.2 State the flame colour given by a potassium ion. [1]

3 Exam-style questions

3.1 A white precipitate forms when dilute nitric acid and silver nitrate are added. The ion present is: [1]

- | | |
|---|---|
| A | B |
| C | D |
- A** sulfate
B carbonate
C chloride
D nitrate

3.2 A solution gives a light blue precipitate with sodium hydroxide that dissolves in excess aqueous ammonia to a deep blue solution.

(a) Name the cation present. [1]

(b) Describe the test for a sulfate ion and its result. [2]

3.3 Describe the test for hydrogen gas and its positive result. [2]

3.4 An unknown white solid **S** dissolves in water. Three tests are carried out.

- Adding dilute nitric acid then silver nitrate gives a **cream** precipitate.
- Adding aqueous sodium hydroxide and warming gives a gas that turns damp red litmus **blue**.
- Adding dilute hydrochloric acid gives **no** effervescence.

(a) Identify the anion and the cation present in **S**, and name the compound. [3]

(b) Name the gas produced in the second test and give the equation for its formation from the ion present. [2]

(c) Explain what the third test rules out, and why. [2]

(d) Describe the tests for carbon dioxide, hydrogen and oxygen, giving the positive result in each case. [3]

(e) Explain why the nitric acid is added before the silver nitrate. [2]
