

# 12.2 Acid-base titrations

Name: \_\_\_\_\_ Class: \_\_\_\_\_ Date: \_\_\_\_\_

Total: 7 marks

**KEY WORDS** indicator 指示剂 • end-point 终点 • titration 滴定 • acid-base titrations 酸碱滴定  
• solution 溶液

## Objective

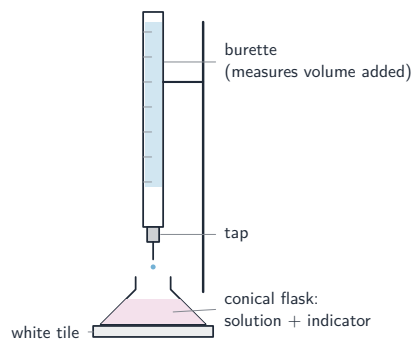
Build the skills to answer exam questions on **acid-base titrations** 酸碱滴定 — the method, the **end-point** 终点, and reading the burette.

You must be able to:

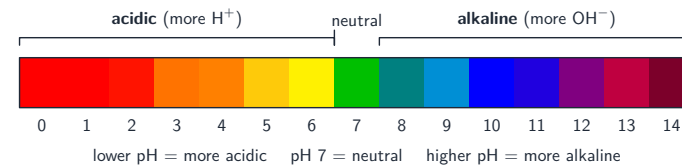
- describe the titration method (pipette, indicator, burette)
- explain the end-point
- relate the procedure to an accurate result

## 1 Worked examples

### ■ The titration method



*Run solution from the burette into the flask until the indicator just changes colour*



*A titration tracks the change in pH*

- pipette a fixed volume of one solution into a flask; add indicator.
- add the other solution from the burette slowly until the colour just changes (end-point).
- read the burette to find the volume added.

## 2 Practice

**2.1** Name the apparatus used to measure the fixed volume into the flask. [1]

**2.2** State what the end-point of a titration shows. [1]

## 3 Exam-style questions

**3.1** The solution added from the burette is run in: [1]

- |   |   |
|---|---|
| A | B |
| C | D |
- A all at once  
B slowly, until the indicator just changes colour  
C until the flask overflows  
D without an indicator

**3.2** A student titrates sodium hydroxide against hydrochloric acid.

(a) Describe how they would carry out the titration.

[3]

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(b) Explain why a few drops of indicator are added.

[1]

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