IGCSE CHEMISTRY • EXERCISE SHEET

5.1 Exothermic and endothermic reactions

Name: _____ Class: _____ Date: _____ **Total: 8 marks**

Objective

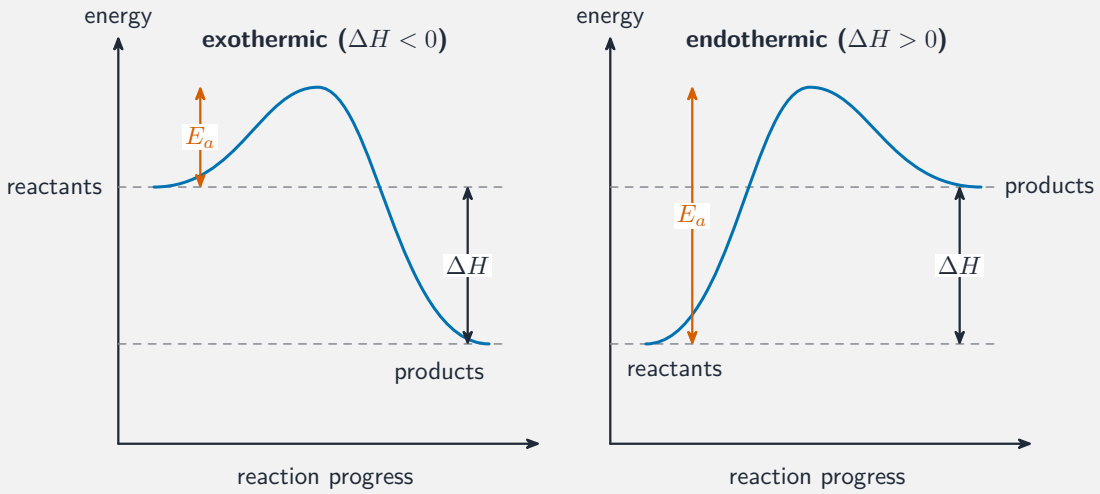
Build the skills to answer exam questions on **exothermic and endothermic reactions** 放热与吸热反应—temperature changes, **enthalpy change** 焓变 ΔH, **reaction pathway diagrams** 反应进程图, and **bond energy** 键能 calculations.

You must be able to:

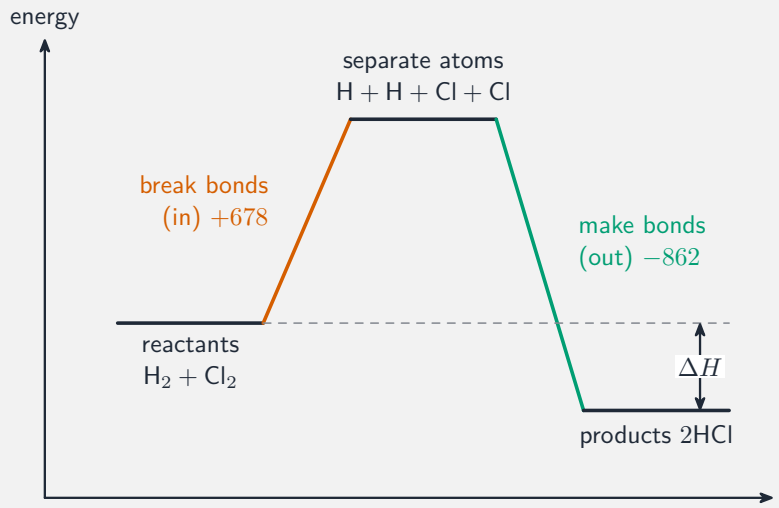
- classify reactions as exothermic ($\Delta H < 0$) or endothermic ($\Delta H > 0$)
- interpret a reaction pathway diagram (activation energy, ΔH)
- calculate ΔH from bond energies

1 Worked examples

Bond energy calculation



Exothermic: products lower ($\Delta H < 0$). Endothermic: products higher ($\Delta H > 0$)



net: $\Delta H = 678 - 862 = -184$ kJ/mol (exothermic)

$$\Delta H = (\textit{bonds broken}) - (\textit{bonds made})$$

$H_2 + Cl_2 \rightarrow 2HCl$: broken $436 + 242 = 678$; made $2 \times 431 = 862$; $\Delta H = 678 - 862 = -184$ kJ/mol.

2 Practice

2.1 State the sign of ΔH for an exothermic reaction. [1]

2.2 State whether bond breaking takes in or gives out energy. [1]

3 Exam-style questions

3.1 During an endothermic reaction, the temperature of the surroundings: [1]

- **A** rises
- **B** falls
- **C** stays the same
- **D** rises then falls

3.2 Consider $H_2 + F_2 \rightarrow 2HF$. Bond energies (kJ/mol): H–H = 436, F–F = 158, H–F = 568.

(a) Calculate ΔH. [3]

(b) State whether the reaction is exothermic or endothermic. [1]

3.3 On a reaction pathway diagram, state what the height of the 'hill' above the reactants represents. [1]

4 Go further

You are now ready for the real exam questions on this subtopic. Open the **5.1 Exothermic and endothermic reactions** past-paper sheet in the Library, or try these in **Practice mode**:

- 0620/21 N25 —Q14 (energy changes)
- 0620/21 N25 —Q15 (bond energy / pathway)

Solutions

2.1 negative.

2.2 it takes in energy (endothermic).

3.1 B. An endothermic reaction takes in energy, so the surroundings cool down.

3.2 (a) bonds broken $= 436 + 158 = 594$; bonds made $= 2 \times 568 = 1136$; $\Delta H = 594 - 1136 = -542$ kJ/mol.
(b) exothermic (ΔH is negative).

3.3 the activation energy (the minimum energy needed for the particles to react).