

IGCSE Chemistry

Name: _____ Score: _____

1. Metals conduct electricity because they have:
 - ☐ free (delocalised) electrons
 - ☐ no electrons
 - ☐ covalent bonds
 - ☐ free ions
2. Copper is used for electrical wiring because it is an excellent:
 - ☐ fuel
 - ☐ magnet
 - ☐ insulator
 - ☐ conductor of electricity
3. An alloy is:
 - ☐ a pure metal
 - ☐ a mixture of a metal with other elements
 - ☐ a type of acid
 - ☐ a non-metal
4. Which of these metals is the most reactive?
 - ☐ iron
 - ☐ copper
 - ☐ potassium
 - ☐ gold
5. Rusting of iron requires which two things? (Select all that apply.)
 - ☐ water
 - ☐ oxygen
 - ☐ nitrogen
 - ☐ sunlight

Tick every correct option.

6. A rock containing enough of a metal compound to extract is called an:
 - ☐ isotope

- ☐ acid
- ☐ alloy
- ☐ ore

7. Which are typical properties of metals? (Select all that apply.)

- ☐ poor electrical conductors
- ☐ malleable
- ☐ shiny
- ☐ good heat conductors

Tick every correct option.

8. Iron is strong but a drawback is that it rusts.

- ☐ True ☐ False

9. Stainless steel contains carbon, which stops it from rusting.

- ☐ True ☐ False

10. Being very unreactive, _____ is found near the bottom of the reactivity series.

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1. **free (delocalised) electrons**

Mobile electrons carry the current.

2. **conductor of electricity**

Copper conducts electricity very well.

3. **a mixture of a metal with other elements**

Alloys are mixtures based on a metal.

4. **potassium**

Potassium is near the top of the reactivity series.

5. **water; oxygen**

Both water and oxygen are needed.

6. **ore**

Metals are extracted from their ores.

7. **malleable; shiny; good heat conductors**

Metals are good electrical conductors, not poor ones.

8. **True**

Rusting is a downside of using iron.

9. **False**

False. Chromium (and often nickel) gives stainless steel its resistance to corrosion. Carbon makes steel harder but does not stop rusting.

10. **gold**

Gold is one of the least reactive metals.