The Periodic Table

IGCSE Chemistry · Vocabulary in context

Word	Example
Periodic Table /ˌpɪərɪˈɒdɪk ˈteɪbl/	The **Periodic Table** lists all the elements.
elements /ˈelɪmənts/	Carbon and oxygen are **elements**.
proton number /ˈprəʊtɒn ˈnʌmbə/	Carbon has a **proton number** of 6. 6
atomic number /əˈtɒmɪk ˈnʌmbə/	Carbon has an **atomic number** of 6. 6
periods /ˈpɪərɪədz/	The rows of the Periodic Table are called **periods**.
groups /gruːps/	The columns of the Periodic Table are called **groups**.
metals /ˈmetlz/	**Metals** form positive ions.
non-metals /nɒn ˈmetlz/	**Non-metals** often form negative ions.
ions /ˈaɪɒnz/	Water molecules surround the **ions** as they dissolve.
electrons /ɪˈlektrɒnz/	**Electrons** move around the nucleus.
electronic configuration /ɪlekˈtrɒnɪk kənˌfɪgjəˈreɪʃn/	The **electronic configuration** of sodium is 2,8,1. 2,8,1
alkali metals /ˈælkəlaɪ ˈmetlz/	The **alkali metals** are in Group 1. 1
lithium /ˈlɪθɪəm/	**Lithium** floats on water and fizzes slowly.
sodium /ˈsəʊdɪəm/	**Sodium** melts into a ball when it reacts with water.

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potassium /pə'tæsɪəm/	**Potassium** reacts with water and burns with a lilac flame.
soft /sɒft/	Sodium is so **soft** that you can cut it with a knife.
trends /trendz/	Hydroxides and sulfates show opposite solubility **trends**.
melting point /'meltɪŋ pɔɪnt/	The **melting point** of ice is 0 °C. 0 °C
density /'densɪti/	Iron has a higher **density** than wood, so it sinks in water.
reactivity /rɪək'tɪvɪti/	The **reactivity** of Group 1 metals increases down the group. 1
halogens /'hælədʒnz/	Chlorine, bromine and iodine are **halogens**. `
chlorine /'klɔ:ri:n/	**Chlorine** gas bleaches damp litmus paper.
bromine /'brəʊmaɪn/	**Bromine** is a red-brown liquid.
iodine /'aɪədi:n/	**Iodine** is a dark grey solid in Group 7. 7
diatomic /ˌdaɪə'tɒmɪk/	Chlorine is **diatomic**: each molecule has two atoms, Cl_2. Cl_2
halide /'hælaɪd/	Lead bromide is a **halide**: a compound of a halogen.
displacement reaction /dɪ'spleɪsmənt rɪ'ækʃn/	In a **displacement reaction**, chlorine pushes bromine out of a compound.
transition elements /træn'sɪʃn 'elɪmənts/	Iron and copper are **transition elements**.

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coloured /ˈkʌləd/	Transition metal compounds are often **coloured**, such as blue copper sulfate.
compounds /ˈkɒmpaʊndz/	Water and salt are **compounds**.
variable /ˈveərɪəbl/	Iron has **variable** charges: Fe^{2+} and Fe^{3+}.
oxidation numbers /ˌɒksɪˈdeɪʃn ˈnʌmbəz/	**Oxidation numbers** show how many electrons an atom has lost or gained.
iron /ˈaɪən/	**Iron** rusts when it meets water and oxygen.
noble gases /ˈnəʊbl ˈgæsɪz/	Helium and neon are **noble gases**: they hardly react.
unreactive /ʌnrɪˈæktɪv/	The noble gases are **unreactive**.
monatomic /ˌmɒnəˈtɒmɪk/	Helium is a **monatomic** gas: its particles are single atoms.
effervescence /ˌefəˈvesəns/	**Effervescence** means lots of bubbles of gas.

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