

17.2 Radioactive decay (full E.3, SL)

Name: _____ Class: _____ Date: _____

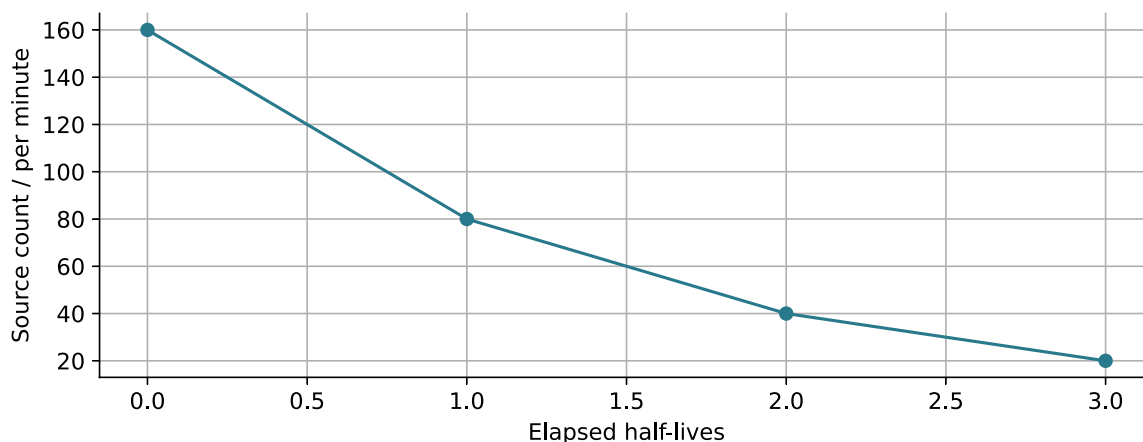
Total: 19 marks

KEY WORDS isotope • binding energy • half-life
• background radiation

Objective

Practise the full guide E.3 toolkit at standard level: isotopes, mass defect and binding energy with $E = \Delta mc^2$, the binding-energy-per-nucleon curve, the strong nuclear force, the random spontaneous nature of decay, alpha, beta-minus, beta-plus and gamma decay equations with neutrinos, penetration and ionizing ability, activity, count rate, background radiation and integer half-lives. Original practice; not an official IB paper.

1 Key knowledge and worked examples



■ The E.3 rule set (SL)

$$E = \Delta mc^2 \quad N = \frac{N_0}{2^n} \quad (n \text{ half-lives})$$

Alpha: $A - 4$, $Z - 2$. Beta-minus: $Z + 1$ with an electron and antineutrino. Beta-plus: $Z - 1$ with a positron and neutrino. Gamma: A , Z unchanged. Penetration falls and ionizing power rises in the order gamma, beta, alpha; subtract background from any count rate.

■ Worked example: counting through background

A sample gives 5200 counts/min with a 200 counts/min background: the source alone is 5000. After three half-lives it contributes $5000/8 = 625$, and the meter reads 825. Two habits: strip the background before halving anything, and check decay equations conserve both A and Z .

2 Practice

2.1 Define an isotope, and state what isotopes of an element share. [2]

2.2 State what is meant by the mass defect of a nucleus. [1]

2.3 Write the full nuclear equation for beta-minus decay of ${}^{14}_6\text{C}$. [2]

2.4 Place alpha, beta and gamma radiation in order of increasing penetrating ability. [1]

3 Exam-style questions

3.1 State the nuclear changes in A and Z for alpha decay, and write the equation for ${}^{226}_{88}\text{Ra}$ decaying by alpha emission. [3]

3.2 Two protons and two neutrons have total mass 4.0319 u ; a helium-4 nucleus has mass 4.0026 u . Calculate the binding energy in MeV ($1\text{ u} = 931.5\text{ MeV}/c^2$) and the binding energy per nucleon. [4]

3.3 Explain why a medical tracer isotope is chosen with a short half-life and gamma emission, referring to penetration and ionizing ability. [3]

3.4 A source reads 4200 counts/min above a background of 200 counts/min . Calculate the count rate after two half-lives, and state one assumption. [3]
