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|-----------------------------------------------|-------------------------------------------|-------------------------------------------|
| <div>T1</div> <div>absolute temperature</div> | <div>T1</div> <div>ideal gas</div>        | <div>T2</div> <div>ionic bond</div>       |
| <div>T2</div> <div>delocalized electron</div> | <div>T4</div> <div>limiting reagent</div> | <div>T4</div> <div>mole</div>             |
| <div>T11</div> <div>functional group</div>    | <div>T11</div> <div>isomer</div>          | <div>T12</div> <div>enthalpy change</div> |

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| <p>T2</p> <p>The electrostatic attraction between oppositely charged ions.</p>                                                   | <p>T1</p> <p>A model of a gas with point particles that exert no intermolecular forces except during perfectly elastic collisions.</p> | <p>T1</p> <p>Temperature on the kelvin scale.</p>                                                                     |
| <p>T4</p> <p>The SI unit of amount of substance, containing exactly 6.02214076 times ten to the power 23 specified entities.</p> | <p>T4</p> <p>The reactant that is used up first and therefore limits the amount of product formed.</p>                                 | <p>T2</p> <p>An electron distributed over several atoms rather than confined to one atom or a single bond.</p>        |
| <p>T12</p> <p>The heat transferred in a reaction or process at constant pressure.</p>                                            | <p>T11</p> <p>A compound with the same molecular formula as another compound but a different arrangement of atoms.</p>                 | <p>T11</p> <p>An atom or group of atoms responsible for characteristic chemical reactions of an organic compound.</p> |

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|-------------------------------------|-----------------------------------------|-----------------------------------|
| <p>T12</p> <p>exothermic</p>        | <p>T14</p> <p>life-cycle assessment</p> | <p>T14</p> <p>functional unit</p> |
| <p>T15</p> <p>titre</p>             | <p>T15</p> <p>equivalence point</p>     | <p>T16</p> <p>rate</p>            |
| <p>T16</p> <p>activation energy</p> | <p>T17</p> <p>reversible reaction</p>   | <p>T17</p> <p>equilibrium</p>     |

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|-----------------------------------------------------------------------------------------------------------------------------------------|------------------------------------------------------------------------------------------------------------------------|----------------------------------------------------------------------------------------------------------|
| <p>T14</p> <p>A specified amount of service or function used as the common basis for comparing products in a life-cycle assessment.</p> | <p>T14</p> <p>Assessment across production, use and disposal.</p>                                                      | <p>T12</p> <p>Describing a process that transfers heat to its surroundings.</p>                          |
| <p>T16</p> <p>A quantity measured or calculated per unit of another quantity, often per unit time.</p>                                  | <p>T15</p> <p>The stage of a titration at which the reactants have been mixed in their stoichiometric proportions.</p> | <p>T15</p> <p>The difference between final and initial readings, and both readings have uncertainty.</p> |
| <p>T17</p> <p>A state with no resultant force or resultant moment.</p>                                                                  | <p>T17</p> <p>A reaction whose products can react to form the original reactants under suitable conditions.</p>        | <p>T16</p> <p>The minimum energy that reacting particles need for a successful reaction.</p>             |

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|---------------------------------------|-------------------------------------|--------------------------------------------|
| <div>T19</div> <div>oxidation</div>   | <div>T19</div> <div>reduction</div> | <div>T22</div> <div>systematic error</div> |
| <div>T22</div> <div>uncertainty</div> |                                     |                                            |
|                                       |                                     |                                            |

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|-------------------------------------------------------------------------------------|-----------------------------------------------------------------------------------------------------|------------------------------------------------------------------------------------------------------|
| <p>T22</p> <p>A consistent measurement bias caused by the method or instrument.</p> | <p>T19</p> <p>A chemical change involving a gain of electrons or a decrease in oxidation state.</p> | <p>T19</p> <p>A chemical change involving a loss of electrons or an increase in oxidation state.</p> |
|                                                                                     |                                                                                                     | <p>T22</p> <p>An estimate of the possible range of error in a measurement.</p>                       |
|                                                                                     |                                                                                                     |                                                                                                      |