

Solutions

2.1 A Calculate the known amount first, apply the stoichiometric ratio, then divide by the unknown solution volume in cubic decimetres. Use concordant titres as required by the school method and report the accepted values.

2.2 titre=final–initial=23.65–1.40=22.25 cm³.

2.3 False. Adding distilled water to the flask changes volume but not the transferred amount of analyte. Do not average a rough titre with carefully measured concordant values.

2.4 titre. titre means The volume delivered between two burette readings.

3.1 $n=0.100 \times 0.0180=0.00180$ mol; $c=0.00180/0.0250=0.0720$ mol/dm³. Overshoot increases recorded volume and inferred concentration in this model. Use concordant runs.