# Electron transfer reactions

International Baccalaureate · IB Diploma · Chemistry · SL · Vocabulary in context

| Word | Example |
|---|---|
| **oxidation** /ˌɒksɪˈdeɪʃn/ | **Oxidation** is the loss of electrons. |
| **reduction** /rɪˈdʌkʃn/ | **Reduction** is the gain of electrons. |
| **oxidation state** /ˌɒksɪˈdeɪʃn steɪt/ | Iron has an **oxidation state** of +3 in Fe2O3.<br>Fe2O3 +3 |
| **oxidizing agent** /ˈɒksɪdaɪzɪŋ ˈeɪdʒənt/ | An **oxidizing agent** gains electrons from another substance. |
| **reducing agent** /rɪˈdjuːsɪŋ ˈeɪdʒənt/ | Carbon is a **reducing agent** in a blast furnace. |
| **half-equation** /hɑːf ɪˈkweɪʒn/ | The **half-equation** Cu2+ + 2e− → Cu shows reduction.<br>Cu2+ + 2e− → Cu |
| **activity series** | The **activity series** puts potassium at the top. |
| **displacement** /dɪˈspleɪsmənt/ | In a **displacement** reaction, chlorine pushes bromine out of a salt. |
| **primary cell** | A **primary cell** is designed for one use. |
| **anode** /ˈænəʊd/ | The **anode** is the positive electrode. |
| **cathode** /ˈkæθəʊd/ | Electrons leave the hot **cathode**. |
| **salt bridge** /sɒlt brɪdʒ/ | A **salt bridge** joins the two half-cells. |
| **secondary cell** | A **secondary cell** can be recharged. |



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| **fuel cell** /ˈfjuːəl sel/ | A **fuel cell** produces electricity while fuel and oxidant are supplied. |
| **electrolytic cell** /ɪˌlektrəˈlɪtɪk sel/ | An **electrolytic cell** uses electricity to drive a chemical change. |
| **electrolysis** /ɪlekˈtrɒləsɪs/ | **Electrolysis** of brine makes chlorine. |
| **electrolyte** /ɪˈlektrəlaɪt/ | Salt water is an **electrolyte**: ions carry the current. |
| **distillation** /dɪstɪˈleɪʃn/ | **Distillation** separates ethanol from water. |
| **reflux** /ˈriːflʌks/ | Heat under **reflux** so that nothing boils away. |
| **hydrogenation** /haɪˈdrɒdʒəneɪʃn/ | **Hydrogenation** of vegetable oil makes margarine. |



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