Structure 3.1 · The periodic table: classification of elements

International Baccalaureate · IB Diploma · Chemistry · SL · Vocabulary in context

Word	Example
periodic table /ˌpɪərɪˈɒdɪk ˈteɪbl/	Find the element's atomic number in the **periodic table**.
periodicity /ˌpɪərɪɒˈdɪsɪti/	**Periodicity** means patterns repeat across each period.
period /ˈpɪərɪəd/	The **period** of the wave is 0.02 s. 0.02 s
group /gruːp/	Sodium and potassium are in the same **group**.
block /blɒk/	Sodium belongs to the s **block** of the table. s
metalloid	Silicon is a **metalloid**.
alkali metal /ˈælkəlaɪ ˈmetl/	Sodium is an **alkali metal**.
halogen /ˈhælədʒn/	Chlorine is a **halogen**.
nuclear charge /ˈnjuːklɪə tʃɑːdʒ/	Chlorine has a larger **nuclear charge** than sodium.
shielding /ˈʃiːldɪŋ/	Inner electrons give **shielding**: they block some of the pull of the nucleus.
electron affinity /ɪˈlektrɒn əˈfɪnɪti/	Chlorine has a negative first **electron affinity**: energy is released.
basic /ˈbeɪsɪk/	A **basic** oxide reacts with an acid.
acidic /əˈsɪdɪk/	Phenol is more **acidic** than ethanol.

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amphoteric /ˌæmfəʊˈterɪk/	An **amphoteric** oxide can react with both acids and bases.
acid deposition	**Acid deposition** damages forests and lakes.
ocean acidification /ˈəʊʃn əˌsɪdɪfɪˈkeɪʃn/	Carbon dioxide dissolving in the sea causes **ocean acidification**.
oxidation state /ˌɒksɪˈdeɪʃn steɪt/	Iron has an **oxidation state** of +3 in Fe2O3. Fe2O3 +3