The ionic model

International Baccalaureate · IB Diploma · Chemistry · SL · Vocabulary in context

Word	Example
ionic model /aɪˈɒnɪk ˈmɒdl/	The **ionic model** explains why salt crystals break.
ion /ˈaɪɒn/	An **ion** is an atom that has gained or lost electrons.
cation /ˈkætaɪən/	Sodium loses one electron to form a **cation**.
anion /ˈænaɪən/	Chlorine gains one electron to form an **anion**.
transition element /trænˈsɪʃn ˈelɪmənt/	Iron is a **transition element**.
polyatomic ion	Sulfate, SO4 2−, is a **polyatomic ion**. SO4 2−
ionic bond /aɪˈɒnɪk bɒnd/	An **ionic bond** is the attraction between oppositely charged ions.
electrostatic attraction /ɪˌlektrəʊˈstætɪk əˈtrækʃn/	**Electrostatic attraction** holds positive and negative ions together.
electronegativity /ɪˌlektrəʊŋgəˈtɪvɪti/	Fluorine has the highest **electronegativity** of all elements.
lattice /ˈlætɪs/	The ions in salt sit in a regular **lattice**.
lattice enthalpy	Magnesium oxide has a large **lattice enthalpy**.
volatility /ˌvɒləˈtɪlɪti/	Bromine has a higher **volatility** than iodine.

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