Ideal gases

International Baccalaureate · IB Diploma · Chemistry · SL ·
Vocabulary in context

Word	Example
kinetic molecular theory /kɪˈnetɪk məˈlekjʊlə ˈθɪəri/	**Kinetic molecular theory** says gas particles are always moving.
ideal gas /aɪˈdɪəl gæs/	An **ideal gas** follows pV = nRT exactly. pV = nRT
intermolecular forces /ˌɪntəməˈlekjʊlə ˈfɔːsɪz/	Weak **intermolecular forces** let a gas boil at a low temperature.
elastic /ɪˈlæstɪk/	In an **elastic** collision, no kinetic energy is lost.
pressure /ˈpreʃə/	A sharp knife puts a large **pressure** on the bread.
kelvin scale /ˈkelvɪn skeɪl/	An ice bath reads about 273 on the **kelvin scale**. 273
absolute zero /ˈæbsəluːt ˈzɪərəʊ/	Nothing can be colder than **absolute zero**.
Avogadro's law	**Avogadro's law** says equal volumes of gas hold equal numbers of particles.
molar volume	The **molar volume** of a gas at room temperature is about 24 dm^3. 24 dm^3
ideal gas equation /aɪˈdɪəl gæs ɪˈkweɪʒn/	The **ideal gas equation** links pressure, volume, temperature and amount. ` `
gas constant /gæs ˈkɒnstənt/	R is the **gas constant**. R

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real gas /rɪəl gæs/

At high pressure, a **real gas** does not follow pV = nRT exactly.

pV = nRT

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