Counting particles by mass: the mole

International Baccalaureate · IB Diploma · Chemistry · SL · Vocabulary in context

Word	Example
mole /məʊl/	One **mole** of carbon has a mass of 12 g. 12 g
amount of substance /əˈmaʊnt ɒv ˈsʌbstəns/	The **amount of substance** is measured in moles.
Avogadro constant /ˌævəˈgædrəʊ ˈkɒnstənt/	The **Avogadro constant** is 6.02×10^{23} per mole. 6.02×10^{23}
formula unit	NaCl is one **formula unit** of sodium chloride. NaCl
relative atomic mass /ˈrelətɪv əˈtɒmɪk mæs/	Chlorine has a **relative atomic mass** of 35.5. 35.5
relative formula mass /ˈrelətɪv ˈfɔːmjʊlə mæs/	Sodium chloride has a **relative formula mass** of 58.5. 58.5
molar mass /ˈməʊlə mæs/	The **molar mass** of water is 18 g/mol. 18 g/mol
empirical formula /emˈpɪrɪkl ˈfɔːmjʊlə/	The **empirical formula** of glucose is CH2O. CH2O
molecular formula /məˈlekjʊlə ˈfɔːmjʊlə/	The **molecular formula** of glucose is C6H12O6. C6H12O6
solution /səˈluːʃn/	Salt water is a **solution**.
solute /ˈsɒljuːt/	Salt is the **solute** in salt water.
solvent /ˈsɒlvənt/	Water is the **solvent** in salt water.

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concentration /ˌkɒnsənˈtreɪʃn/	The **concentration** of the acid is 2 mol/dm^3. 2 mol/dm^3
standard solution /ˈstændəd səˈluːʃn/	A **standard solution** has a known concentration.
Avogadro's law	**Avogadro's law** says equal volumes of gas hold equal numbers of particles.