Introduction to the particulate nature of matter

International Baccalaureate · IB Diploma · Chemistry · SL · Vocabulary in context

Word	Example
• **element** /'elɪmənt/	Each **element** of the array holds one value.
compound /'kɒmpaʊnd/	Water is a **compound** of hydrogen and oxygen.
mixture /'mɪkstʃə/	Air is a **mixture** of gases.
homogeneous /ˌhɒmə'dʒi:niəs/	Salt water is a **homogeneous** mixture: it looks the same all through.
heterogeneous	Sand in water makes a **heterogeneous** mixture.
filtration /fɪl'treɪʃn/	Use **filtration** to remove sand from water.
evaporation /ɪˌvæpə'reɪʃn/	**Evaporation** dries wet clothes.
crystallization	**Crystallization** gives pure salt crystals from salt water.
distillation /dɪstɪ'leɪʃn/	**Distillation** separates ethanol from water.
solvation	In **solvation**, water molecules surround each dissolved ion.
saturated /'sætʃəreɪtɪd/	A **saturated** solution cannot dissolve any more solid.
kinetic molecular theory /kɪ'netɪk mə'lekjʊlə 'θɪəri/	**Kinetic molecular theory** says gas particles are always moving.

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sublimation /ˌsʌblɪ'meɪʃn/	Dry ice undergoes **sublimation** from solid to gas.
endothermic /ˌendəʊ'θɜ:mɪk/	An **endothermic** change takes in heat.
deposition	Frost forms by **deposition** of water vapour as ice.
exothermic /egzə'ðɜ:mɪk/	An **exothermic** reaction gives out heat.
purity /'pjʊərɪti/	A sharp melting point shows high **purity**.
average kinetic energy /'ævrɪdʒ kɪ'netɪk 'enədʒi/	Hot gas particles have a higher **average kinetic energy**.
absolute zero /'æbsəlu:t 'zɪərəʊ/	Nothing can be colder than **absolute zero**.