Introduction to the particulate nature of matter

International Baccalaureate · IB Diploma · Chemistry · SL · Vocabulary in context

Word	Example
• **element** /ˈelɪmənt/	Each **element** of the array holds one value.
compound /ˈkɒmpaʊnd/	Water is a **compound** of hydrogen and oxygen.
mixture /ˈmɪkstʃə/	Air is a **mixture** of gases.
homogeneous /ˌhɒməˈdʒiːniəs/	Salt water is a **homogeneous** mixture: it looks the same all through.
heterogeneous	Sand in water makes a **heterogeneous** mixture.
filtration /fɪlˈtreɪʃn/	Use **filtration** to remove sand from water.
evaporation /ɪˌvæpəˈreɪʃn/	**Evaporation** dries wet clothes.
crystallization	**Crystallization** gives pure salt crystals from salt water.
distillation /dɪstɪˈleɪʃn/	**Distillation** separates ethanol from water.
solvation	In **solvation**, water molecules surround each dissolved ion.
saturated /ˈsætʃəreɪtɪd/	A **saturated** solution cannot dissolve any more solid.
kinetic molecular theory /kɪˈnetɪk məˈlekjʊlə ˈθɪəri/	**Kinetic molecular theory** says gas particles are always moving.

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sublimation /ˌsʌblɪˈmeɪʃn/	Dry ice undergoes **sublimation** from solid to gas.
endothermic /ˌendəʊˈθɜːmɪk/	An **endothermic** change takes in heat.
deposition	Frost forms by **deposition** of water vapour as ice.
exothermic /egzəˈðɜːmɪk/	An **exothermic** reaction gives out heat.
purity /ˈpjʊərɪti/	A sharp melting point shows high **purity**.
average kinetic energy /ˈævrɪdʒ kɪˈnetɪk ˈenədʒi/	Hot gas particles have a higher **average kinetic energy**.
absolute zero /ˈæbsəluːt ˈzɪərəʊ/	Nothing can be colder than **absolute zero**.

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