10.1 The periodic table and periodic trends

Name: _____ Class: _____ Date: _____ **Total: 30 marks**

KEY WORDS period • group • ionization energy • atomic radius • halogen • oxidation state

Objective

Build the skills to answer exam questions on **the periodic table** : periods, groups and blocks, periodic trends, groups 1 and 17, oxides and oxidation states.

You must be able to:

- deduce the period, group and block of an element from its electron configuration
- explain trends in atomic radius, ionic radius, ionization energy, electron affinity and electronegativity
- describe the reactions of group 1 metals with water and of halogens with halide ions
- write equations for metal and non-metal oxides with water, and explain acid rain and ocean acidification

1 Key knowledge and worked examples

Organizing the elements

The **period** number is the number of occupied energy levels; the **group** shows the outer electrons. The **block** (s, p, d, f) is the sublevel being filled: $[Ne]3s^2 3p^4$ is period 3, group 16, p-block (sulfur).

trend	across a period (→)	down a group (↓)
atomic radius	decreases: more protons, same shielding, electrons pulled closer	increases: more occupied levels
first ionization energy	generally increases	decreases: outer electron further away and more shielded
electronegativity	increases	decreases
metallic character	decreases	increases

Positive ions are **smaller** than their atoms; negative ions are **larger**.

group	behaviour
1 (alkali metals)	$2M + 2H_2O \rightarrow 2MOH + H_2$; more reactive down the group (outer electron lost more easily)
17 (halogens)	a more reactive halogen displaces a less reactive one: $Cl_2 + 2Br^- \rightarrow 2Cl^- + Br_2$; less reactive down the group

Oxides: metal oxides are **basic** ($Na_2O + H_2O \rightarrow 2NaOH$); non-metal oxides are **acidic** ($SO_3 + H_2O \rightarrow H_2SO_4$); Al_2O_3 is amphoteric. SO_2 and NO_x cause **acid rain**; dissolved CO_2 causes **ocean acidification**. In a compound, an **oxidation state** is the charge an atom would have if all bonds were ionic.

Worked example 1: radius. Na^+ (102 pm) is much smaller than Na (186 pm): it has **lost its outer level**.

Worked example 2: oxidation state. In SO_4^{2-}: $x + 4(-2) = -2$, so S is +6.

Answer shapes: nuclear charge, distance and shielding, then the effect on attraction.

2 Practice

2.1 An element has the configuration $[Ne]3s^2 3p^3$. State its period, group and block. [2]



2.2 Explain why atomic radius decreases from sodium to chlorine. [3]

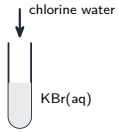


2.3 Explain why the first ionization energy decreases down group 1. [3]

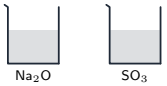


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2.4 Chlorine water is added to potassium bromide solution. State what you would observe, write the ionic equation and explain the reaction. [3]



2.5 Write equations for the reactions of sodium oxide and of sulfur trioxide with water, and state whether each solution is acidic or alkaline. [3]



3 Exam-style questions

3.1 Potassium reacts vigorously with water, with a lilac flame.

(a) Write an equation, with state symbols, for the reaction. [2]

(b) Explain why potassium reacts more vigorously than sodium. [3]

(c) State and explain the pH of the solution formed. [2]

(d) Describe and explain the trend in electronegativity across period 3. [2]

3.2 Burning fossil fuels releases oxides of sulfur and nitrogen. Carbon dioxide dissolves in the oceans.

(a) Deduce the oxidation state of sulfur in SO_2 and of nitrogen in NO_2. [2]

(b) Describe how acid rain forms from sulfur dioxide, and give one effect of acid rain. [3]



(c) Write equations to show how dissolved carbon dioxide makes the oceans more acidic. [2]

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