

T1 absolute temperature	T1 ideal gas	T2 ionic bond
T2 delocalized electron	T4 limiting reagent	T4 mole
T11 functional group	T11 isomer	T12 enthalpy change

T2 The electrostatic attraction between oppositely charged ions.	T1 A model of a gas with point particles that exert no intermolecular forces except during perfectly elastic collisions.	T1 Temperature on the kelvin scale.
T4 The SI unit of amount of substance, containing exactly 6.02214076 times ten to the power 23 specified entities.	T4 The reactant that is used up first and therefore limits the amount of product formed.	T2 An electron distributed over several atoms rather than confined to one atom or a single bond.
T12 The heat transferred in a reaction or process at constant pressure.	T11 A compound with the same molecular formula as another compound but a different arrangement of atoms.	T11 An atom or group of atoms responsible for characteristic chemical reactions of an organic compound.

T12 exothermic	T14 life-cycle assessment	T14 functional unit
T15 entropy	T15 Gibbs energy	T16 titre
T16 equivalence point	T17 rate	T17 activation energy

T14 A specified amount of service or function used as the common basis for comparing products in a life-cycle assessment.	T14 Assessment across production, use and disposal.	T12 Describing a process that transfers heat to its surroundings.
T16 The difference between final and initial readings, and both readings have uncertainty.	T15 A thermodynamic state function equal to enthalpy minus absolute temperature times entropy.	T15 A thermodynamic measure related to energy dispersal and the number of microscopic arrangements consistent with a system's macroscopic state.
T17 The minimum energy that reacting particles need for a successful reaction.	T17 A quantity measured or calculated per unit of another quantity, often per unit time.	T16 The stage of a titration at which the reactants have been mixed in their stoichiometric proportions.

T18 reversible reaction	T18 equilibrium	T20 oxidation
T20 reduction	T23 systematic error	T23 uncertainty

T20 A chemical change involving a loss of electrons or an increase in oxidation state.	T18 A state with no resultant force or resultant moment.	T18 A reaction whose products can react to form the original reactants under suitable conditions.
T23 An estimate of the possible range of error in a measurement.	T23 A consistent measurement bias caused by the method or instrument.	T20 A chemical change involving a gain of electrons or a decrease in oxidation state.