Reactivity 3.1 · Proton transfer reactions

International Baccalaureate · IB Diploma · Chemistry · HL · Vocabulary in context

Word	Example
proton /ˈprəʊtɒn/	A **proton** has a positive charge.
Brønsted–Lowry acid	A **Brønsted–Lowry acid** donates a proton.
Brønsted–Lowry base	A **Brønsted–Lowry base** accepts a proton.
conjugate acid–base pair /ˈkɒndʒuːgeɪt ˈæsɪd beɪs peə/	HCl and Cl− are a **conjugate acid–base pair**. HCl Cl−
amphiprotic	Water is **amphiprotic**: it can give or take a proton.
ionic product of water /aɪˈɒnɪk ˈprɒdʌkt ɒv ˈwɔːtə/	The **ionic product of water** is 1.0×10^{-14} mol^2 dm^{-6} at 25 °C. 25 °C 1.0×10^{-14} mol^2 dm^{-6}
strong acid /strɒŋ ˈæsɪd/	Hydrochloric acid is a **strong acid**.
weak acid /wiːk ˈæsɪd/	Vinegar contains a **weak acid**.
neutralization /ˌnjuːtrəlaɪˈzeɪʃn/	**Neutralization** is a reaction between an acid and a base.
pH curve	A **pH curve** plots pH against added titrant volume. pH
equivalence point /ɪˈkwɪvələns pɔɪnt/	At the **equivalence point**, the acid and base have exactly reacted.
acid dissociation constant /ˈæsɪd dɪˌsəʊsɪˈeɪʃn ˈkɒnstənt/	A weak acid has a small **acid dissociation constant**.

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indicator /ˈɪndɪkeɪtə/	The **indicator** changes colour at the end point.
buffer solution /ˈbʌfə səˈluːʃn/	A **buffer solution** keeps the pH almost the same.



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