How far? The extent of chemical change

International Baccalaureate · IB Diploma · Chemistry · HL · Vocabulary in context

Word	Example
reversible reaction /rɪˈvɜːsɪbl rɪˈækʃn/	A **reversible reaction** can go both ways.
dynamic equilibrium /daɪˈnæmɪk ˌiːkwɪˈlɪbrɪəm/	At **dynamic equilibrium**, both reactions go at the same rate.
closed system /kləʊzd ˈsɪstəm/	In a **closed system**, no energy goes in or out.
isotope /ˈaɪsətəʊp/	Carbon-14 is an **isotope** of carbon. Carbon-14
equilibrium constant /ˌiːkwɪˈlɪbrɪəm ˈkɒnstənt/	A large **equilibrium constant** means mostly products.
position of equilibrium /pəˈzɪʃn ɒv ˌiːkwɪˈlɪbrɪəm/	Changing pressure can shift the **position of equilibrium**.
Le Chatelier's principle /lə ˈtʃeɪtlɪəz ˈprɪnsɪpl/	**Le Chatelier's principle** predicts how an equilibrium moves.
yield /jiːld/	A higher **yield** means more product is obtained.
compromise /ˈkɒmprəmaɪz/	A **compromise** temperature balances reaction speed and product yield.
reaction quotient /rɪˈækʃn ˈkwəʊʃənt/	The **reaction quotient** uses current concentrations to predict reaction direction.
approximation /əˌprɒksɪˈmeɪʃn/	An **approximation** simplifies a calculation when a change is small.

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∑ **Gibbs energy** /gɪbz ˈenədʒi/	A negative change in **Gibbs energy** means the reaction can happen.

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