Reactivity 1.2 · Energy cycles in reactions

International Baccalaureate · IB Diploma · Chemistry · HL · Vocabulary in context

Word	Example
energy cycle /ˈenədʒi ˈsaɪkl/	Draw an **energy cycle** with two routes from the same start to the same end.
bond enthalpy /bɒnd enˈθælpi/	**Bond enthalpy** is the energy needed to break one mole of bonds.
average bond enthalpy	The **average bond enthalpy** of C–H is 414 kJ/mol. C–H 414 kJ/mol
Hess's law /ˈhesɪz lɔː/	**Hess's law** says the enthalpy change is the same by any route.
standard enthalpy change of formation	The **standard enthalpy change of formation** of water is −286 kJ/mol. −286 kJ/mol
standard enthalpy change of combustion	The **standard enthalpy change of combustion** of methane is −890 kJ/mol. −890 kJ/mol
Born–Haber cycle /bɔːn ˈheɪbə ˈsaɪkl/	A **Born–Haber cycle** finds the lattice enthalpy.
enthalpy of atomization	The **enthalpy of atomization** turns the solid element into gas atoms.

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