

2.5 Use atomic numbers, isotopes and bonding models correctly

Name: _____ Class: _____ Date: _____ Total: **24 marks**

KEY WORDS atomic number 原子序数 • mass number 质量数 • isotope 同位素

Objective

Use atomic numbers, isotopes and bonding models correctly. This is original GAC023-aligned practice, not an official assessment. All supplied case data and review records are fictional; explicitly cited background is factual. The case data are invented for learning; they are not findings about a real school, country or population. Use the supplied evidence and do not invent missing observations.

1 Key knowledge and worked examples

■ Key knowledge: particle counts distinguish elements, isotopes and ions

The **atomic number** 原子序数 is the proton count and identifies an element. The **mass number** 质量数 counts protons plus neutrons. An **isotope** 同位素 has the same proton count as another isotope of its element but a different neutron count. A neutral atom has equal proton and electron counts; changing electrons can form an ion without changing the element. Ionic bonding is electrostatic attraction between oppositely charged ions; electron transfer can explain their formation. Covalent bonding involves shared electron pairs.

■ Source A: text or task brief

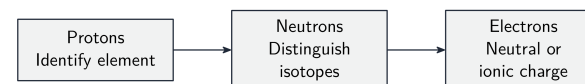
Supplied introductory models: neutral sodium-23 has atomic number 11 and mass number 23; neutral sodium-24 has atomic number 11 and mass number 24. Chlorine-35 has atomic number 17 and mass number 35. In a simplified ionic formation model, one sodium atom loses one electron to form Na^+ and one chlorine atom gains one electron to form Cl^- . The ions attract; the model is not a practical instruction to react these materials.

The periodic table orders elements by atomic number. In the introductory main-group model, outer electron arrangements help explain similarities in group behaviour. For a covalent example, two hydrogen atoms share an electron pair in H_2 . Hydrogen-1 has one proton and no neutron.

■ Source B: response or review record

Student note: “Sodium-23 has 23 neutrons. Sodium-24 is a different element because its mass changed. Na^+ has twelve electrons and Cl^- has sixteen. Ionic bonds are simply the act of transferring an electron, with no attraction afterwards. Every atom must contain a neutron. Elements in a group are identical because all their electrons are the same.”

■ Worked example: count particles before describing charge



Known: sodium-23 has proton number 11 and mass number 23. Why subtract proton number? The mass number counts protons and neutrons together.

$$N = A - Z$$

$$N = 23 - 11 = 12$$

The neutral atom has 11 electrons. Losing one gives Na^+ with 10 electrons, not 12. The positive charge arises because protons outnumber electrons by one; the proton count is unchanged.

2 Practice

2.1 State the meanings of atomic number and mass number. [2]

2.2 Calculate the neutron counts of sodium-23 and sodium-24. [2]

2.3 Give the electron counts of Na^+ and Cl^- and explain one charge. [3]

2.4 Explain why the two sodium isotopes are the same element but not identical atoms.[3]

3 Exam-style questions

3.1 Correct the bonding, hydrogen and group claims using six points. [6]

3.2 Write an 80–110-word corrected particle-and-bonding note. [8]
