

# Radioactivity and particles

Pearson Edexcel · International GCSE · Physics

## The nucleus at work

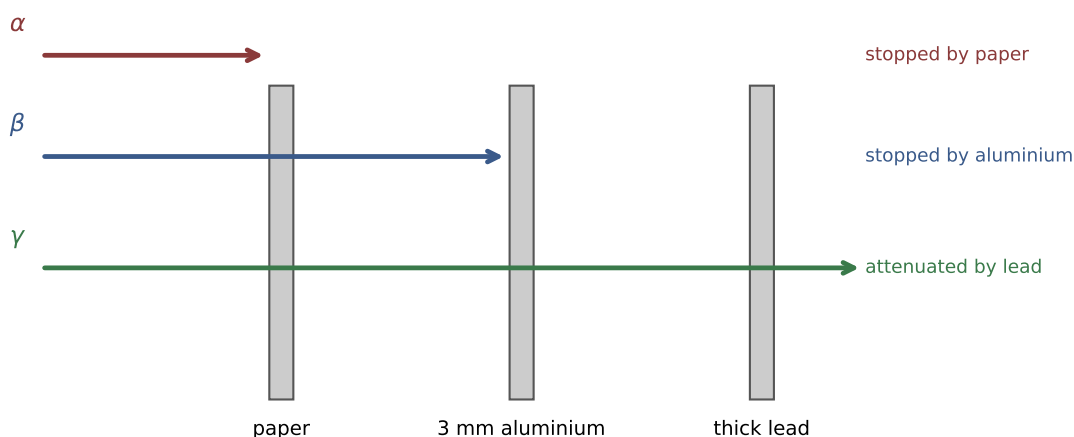
Some nuclei are unstable: they throw out radiation at random, decay by half-lives, and – in reactors and stars – release energy by splitting or joining. This section covers atomic notation, the three ionising radiations with their penetrating powers, half-life calculations, the fission chain reaction, and why fusion lights the stars but not our power stations.

**Edexcel International GCSE Physics 4PH1, Section 7** (statements 7.1–7.26).  
Units: becquerel (Bq), with time in s, min or h.

## Radioactivity

An atom holds protons and neutrons in a nucleus; notation  ${}^A_ZX$  gives **mass (nucleon) number** A and **atomic (proton) number** Z. **Isotopes** share Z but differ in neutrons. Unstable nuclei emit – randomly – **alpha**  $\alpha$  (helium nucleus,  ${}^4_2\text{He}$ ), **beta-minus**  $\beta^-$  (an electron, created as a neutron turns proton) and **gamma**  $\gamma$  (high-energy electromagnetic wave, no particles). Balancing: both **mass number and charge are conserved**.

**Worked check.** Carbon-14:  ${}^{14}_6\text{C} \rightarrow {}^{14}_7\text{N} + {}^0_{-1}\text{e}$  – A stays 14; charge  $6 = 7 + (-1)$  ✓.  
Radon:  ${}^{222}_{86}\text{Rn} \rightarrow {}^{218}_{84}\text{Po} + {}^4_2\text{He}$ .

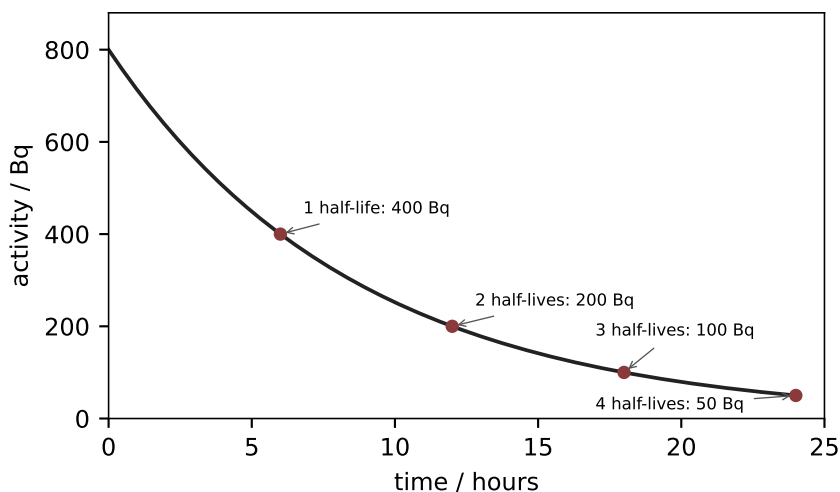


*Alpha stopped by paper, beta by a few mm of aluminium, gamma attenuated by thick lead.*

| Radiation | Nature        | Ionising  | Range in air |
|-----------|---------------|-----------|--------------|
| alpha     | $2p + 2n$     | strongest | a few cm     |
| beta      | fast electron | medium    | tens of cm   |
| gamma     | EM wave       | weakest   | metres       |

Detection: photographic film (badge) or a **Geiger–Müller tube**. **Background radiation** comes from rocks and cosmic rays (space), food, medical uses and fallout.

**Activity** (in becquerels;  $1 \text{ Bq} = 1 \text{ decay/s}$ ) falls steadily: the **half-life** is the time for half the undecayed nuclei to decay. **Worked check.** 800 Bq with half-life 6 h: after 24 h (4 half-lives) activity =  $800/2^4 = 50 \text{ Bq}$ .

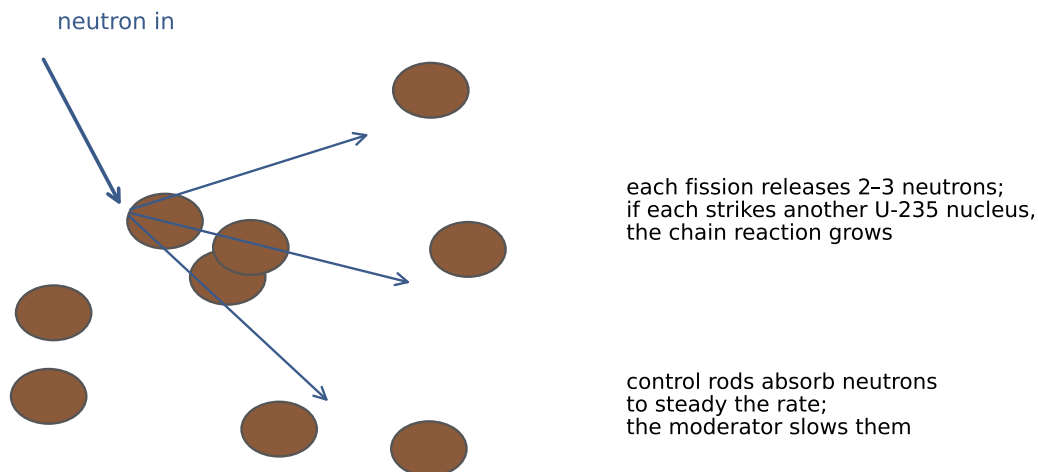


*A decay curve halving every half-life, annotated at each halving.*

Uses: medical tracers (short half-life, gamma), cancer treatment, sterilising, thickness control (beta through paper). **Contamination** (radioactive material on you – must be removed) differs from **irradiation** (exposure to radiation – stops when the source leaves). Dangers: mutation of DNA, cell and tissue damage; radioactive waste must be stored safely for its long half-lives.

## Fission and fusion

Split a big nucleus or join two small ones – both release energy, and each powers a different machine.



*Each fission releases neutrons that can strike further U-235 nuclei; control rods and moderator manage it.*

**Fission** : a neutron strikes **U-235**, splitting it into **two radioactive daughter nuclei**, 2-3 neutrons and **kinetic energy** – and if those neutrons strike more U-235, a **chain reaction** runs. The **moderator** slows neutrons so they split nuclei efficiently; **control rods** absorb neutrons to hold the rate steady; **shielding** protects workers. **Fusion** joins small nuclei into a larger one, converting mass to energy – the Sun's source. Fusion needs enormous temperature and pressure because **positively charged protons electrostatically repel** – the nuclei must collide fast enough to overcome that repulsion.

## Check yourself

1. In  ${}^{19}_9F$ , state the numbers of protons, neutrons and electrons in the neutral atom.
2. Name the three ionising radiations and their natures.
3. Write the full equation for alpha decay of  ${}^{226}_{88}Ra$ .
4. Which radiation is stopped by paper, and which passes through aluminium?
5. Give two sources of background radiation.
6. Define half-life.
7. A source of 3200 Bq has half-life 8 days. Find its activity after 32 days.
8. State the difference between contamination and irradiation.
9. Name the roles of the moderator, control rods and shielding in a reactor.
10. Why does fusion need such high temperatures?

Answers: 1 9 protons, 10 neutrons, 9 electrons; 2 alpha – helium nucleus (2p+2n); beta-minus – fast electron; gamma – high-energy electromagnetic wave; 3  ${}^{226}_{88}Ra \rightarrow$

${}_{86}^{222}\text{Rn} + {}_2^4\text{He}$ ; 4 alpha stopped by paper; gamma passes aluminium (needs thick lead); 5 any two: rocks/building materials, cosmic rays, food/drink, medical, radon gas, nuclear fallout; 6 the time taken for half the undecayed nuclei (or the activity) to decay; 7 32 days = 4 half-lives:  $3200 \rightarrow 1600 \rightarrow 800 \rightarrow 400 \rightarrow 200$  Bq; 8 contamination – radioactive material stuck to you, hazard remains until removed; irradiation – exposure to radiation from a source, ends when the source is removed; 9 moderator slows neutrons for efficient fission; control rods absorb neutrons to control the rate; shielding absorbs escaping radiation to protect workers; 10 the positively charged protons of two nuclei repel electrostatically, so nuclei must collide at enormous speed – only reached at very high temperature and pressure – to get close enough to fuse.