T1 melting	T1 freezing	T1 boiling
T1 evaporation	T1 condensation	T1 sublimation
T1 solvent	T1 solute	T1 solution

T1 Vaporisation throughout a liquid when its vapour pressure matches the surrounding pressure.	T1 The change of state from liquid to solid.	T1 The change of a substance from solid to liquid.
T1 The direct change of a substance from solid to gas without passing through a liquid phase.	T1 The change of a substance from a gas to a liquid.	T1 The escape of particles from the surface of a liquid into its vapour without the whole liquid boiling.
T1 A homogeneous mixture in which one or more solutes are dissolved in a solvent.	T1 A substance dissolved in a solvent to form a solution.	T1 The substance in which a solute dissolves to form a solution.

T1 saturated solution	T1 solubility curve	T1 element
T1 compound	T1 mixture	T1 filtration
T1 crystallisation	T1 simple distillation	T1 fractional distillation

T1 A substance made of atoms with the same proton number.	T1 A graph showing the maximum mass of solute dissolving in a stated mass of solvent at different temperatures, often reported per 100 g of solvent.	T1 A solution containing the solute's equilibrium solubility limit at the stated conditions.
T1 Separating insoluble particles from a fluid by passing the mixture through a porous barrier.	T1 A combination of substances that are not chemically bonded together in fixed proportions.	T1 A substance containing two or more different elements chemically bonded in fixed proportions.
T1 Separating components of a liquid mixture through repeated evaporation and condensation based on differences in volatility.	T1 Boiling off and condensing the solvent to separate it from a dissolved solid.	T1 Growing crystals from a cooling or evaporating saturated solution to recover a dissolved solid.

T1 paper chromatography	T1 Atomic number	T1 mass number
T1 Isotope	T1 group	T1 period
T1 state symbol	T1 mole	T1 Percentage yield

T1 The total number of protons and neutrons in an atomic nucleus.	T1 Protons.	T1 Separating dissolved substances by their different movement with a solvent through paper; Rf is spot-centre distance from the start line divided by solvent-front distance from the same line.
T1 A horizontal row of the periodic table, corresponding to the principal shell occupied by the outer electrons of its elements.	T1 A vertical column in the periodic table whose elements have related outer-electron configurations and chemical properties.	T1 Atoms of the same element with the same proton number but different numbers of neutrons.
T1 The actual product mass as a percentage of the theoretical maximum.	T1 The SI unit of amount of substance, containing exactly 6.02214076 times ten to the power 23 specified entities.	T1 A label in a chemical equation identifying a substance as solid, liquid, gas or dissolved in water.

T1 Empirical formula	T1 Ionic bonding	T1 cation
T1 anion	T1 Covalent bonding	T1 Electrolysis
T1 cathode	T1 anode	T1 oxidation

T1 An ion with a positive electric charge.	T1 Metals lose electrons to non-metals.	T1 The simplest whole-number ratio of atoms in a compound.
T1 Using electrical energy to drive reactions in an electrolyte: reduction occurs at the cathode and oxidation at the anode; aqueous products depend on the ions and conditions.	T1 Bonding in which atoms share pairs of electrons, attracted to the nuclei of both bonded atoms.	T1 An ion with a negative electric charge.
T1 A chemical change involving a loss of electrons or an increase in oxidation state.	T1 An electrode at which oxidation occurs.	T1 An electrode at which reduction occurs.

T1 reduction	T2 greenhouse gas	T2 rusting
T2 galvanising	T2 sacrificial protection	T2 oxidation
T2 reduction	T2 oxidising agent	T2 reducing agent

T2 The corrosion of iron, needing both water and oxygen.	T2 A gas that absorbs and emits infrared radiation, contributing to atmospheric warming.	T1 A chemical change involving a gain of electrons or a decrease in oxidation state.
T2 A chemical change involving a loss of electrons or an increase in oxidation state.	T2 Protecting a metal by electrically connecting it to a more reactive metal, which oxidises preferentially.	T2 Coating iron or steel with zinc to protect it from corrosion.
T2 A substance that donates electrons to another substance and is itself oxidised.	T2 A substance that accepts electrons from another substance and is itself reduced.	T2 A chemical change involving a gain of electrons or a decrease in oxidation state.

T2 alloy	T2 titration	T2 pH scale
T2 flame test	T2 precipitate	T3 exothermic
T3 endothermic	T3 calorimetry	T3 catalyst

T2 A logarithmic measure of acidity; in dilute aqueous solutions near 25 °C, pH 7 is neutral, lower values acidic and higher values alkaline. The familiar 0–14 range is not an absolute limit.	T2 A method of finding an unknown concentration by measuring the volume needed to react with a solution of known concentration.	T2 A metallic material combining a metal with other elements, which may include nonmetals; many alloys are harder than the corresponding pure metal.
T3 Describing a process that transfers heat to its surroundings.	T2 An insoluble solid formed when substances in solution react.	T2 Identifying a metal ion by the colour its compound gives in a flame.
T3 A substance that increases a reaction rate without being used up overall.	T3 Measuring heat transfer using the temperature change and known heat capacity of a substance or calorimeter, such as water or a reaction solution.	T3 Describing a process that takes in heat from its surroundings.

T3 collision theory	T3 activation energy	T3 reversible reaction
T3 dynamic equilibrium	T4 hydrocarbon	T4 homologous series
T4 functional group	T4 isomerism	T4 isomer

T3 A reaction whose products can react to form the original reactants under suitable conditions.	T3 The minimum energy that reacting particles need for a successful reaction.	T3 The explanation that reactions require collisions with sufficient energy and suitable orientation.
T4 A family of compounds with the same functional group and successive CH_2 units, such as the alkanes.	T4 A compound of hydrogen and carbon only.	T3 A state in which opposing processes continue at equal rates, so there is no overall change in composition.
T4 One of the compounds sharing a molecular formula but differing in structure.	T4 The existence of compounds with the same molecular formula but different structures.	T4 An atom or group of atoms responsible for characteristic chemical reactions of an organic compound.

T4 displayed formula	T4 fractional distillation	T4 incomplete combustion
T4 cracking	T4 alkane	T4 alkene
T4 saturated	T4 unsaturated	T4 substitution

T4 Burning with insufficient oxygen for complete oxidation, often producing carbon monoxide or carbon as well as other products.	T4 Separating components of a liquid mixture through repeated evaporation and condensation based on differences in volatility.	T4 A formula showing every atom and bond in a molecule.
T4 The open-chain hydrocarbons with one carbon–carbon double bond studied here, with general formula C_nH_{2n}.	T4 The saturated, open-chain hydrocarbons studied here, with only single carbon–carbon bonds and general formula C_nH_{2n+2}.	T4 Breaking long-chain alkanes into smaller hydrocarbons, including alkanes and alkenes; catalytic cracking uses heat and a catalyst.
T4 A reaction in which an atom or group in a molecule is replaced by another atom or group.	T4 Containing one or more carbon-carbon double or triple bonds rather than only single bonds.	T4 Describing a solution unable to dissolve more solute under stated conditions, or an organic molecule with no carbon-carbon double or triple bonds.

T4 addition	T4 alcohol	T4 fermentation
T4 carboxylic acid	T4 ester	T4 ethyl ethanoate
T4 addition polymer	T4 monomer	T4 polyester

T4 The breakdown of sugars without oxygen, often by microorganisms, to release energy and form products such as ethanol or lactic acid.	T4 An organic compound with a hydroxyl group bonded to a saturated carbon atom.	T4 A reaction adding atoms across a C=C double bond, as in hydrogenation or bromination of an alkene.
T4 The ester formed from ethanol and ethanoic acid, with a fruity smell.	T4 An organic compound containing a carbonyl group bonded to an oxygen atom that is bonded to another carbon group.	T4 An organic compound containing the COOH functional group.
T4 A polymer containing repeated ester links in its main chain.	T4 A molecule that can join with other molecules to form a polymer.	T4 A polymer formed by joining alkene monomers with no small molecule lost, as in poly(ethene).