Principles of chemistry

Pearson Edexcel · International GCSE · Chemistry · Vocabulary in context

Word	Example
melting /ˈmeltɪŋ/	**Melting** turns ice into water.
freezing /ˈfriːzɪŋ/	**Freezing** turns water into ice.
boiling /ˈbɔɪlɪŋ/	During **boiling**, bubbles form all through the water.
evaporation /ɪˌvæpəˈreɪʃn/	**Evaporation** dries wet clothes.
condensation /kɒnden'seɪʃn/	**Condensation** forms water droplets as moist air cools.
sublimation /ˌsʌblɪˈmeɪʃn/	Dry ice undergoes **sublimation** from solid to gas.
solvent /ˈsɒlvənt/	Water is the **solvent** in salt water.
solute /ˈsɒljuːt/	Salt is the **solute** in salt water.
solution /səˈluːʃn/	Salt water is a **solution**.
saturated solution /ˈsætʃəreɪtɪd səˈluːʃn/	A **saturated solution** cannot dissolve any more solid.
solubility curve /ˌsɒljuːˈbɪlɪti kɜːv/	Read the **solubility curve** at the temperature stated.
• **element** /ˈelɪmənt/	Each **element** of the array holds one value.
compound /ˈkɒmpaʊnd/	Water is a **compound** of hydrogen and oxygen.
mixture /ˈmɪkstʃə/	Air is a **mixture** of gases.

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filtration /fɪl'treɪʃn/	Use **filtration** to remove sand from water.
crystallisation /ˌkrɪstəlaɪ'zeɪʃn/	**Crystallisation** gives pure salt crystals from a solution.
simple distillation /'sɪmpl dɪstɪ'leɪʃn/	**Simple distillation** gets pure water from salty water.
fractional distillation /'frækʃənl dɪstɪ'leɪʃn/	**Fractional distillation** separates crude oil into fractions.
paper chromatography /'peɪpə krəʊmə'tɒgrəfi/	Coloured spots separate as solvent moves during **paper chromatography**.
Atomic number /ə'tɒmɪk 'nʌmbə/	**Atomic number** eight identifies an oxygen atom.
mass number /mæs 'nʌmbə/	Carbon-12 has a **mass number** of 12. Carbon-12 12
Isotope /'aɪsətəʊp/	Carbon-14 is a radioactive **isotope** of carbon. Carbon-14
group /gru:p/	Sodium and potassium are in the same **group**.
period /'pɪərɪəd/	The **period** of the wave is 0.02 s. 0.02 s
state symbol /steɪt 'sɪmbl/	The **state symbol** (aq) identifies a substance dissolved in water. (aq)
mole /məʊl/	One **mole** of carbon has a mass of 12 g. 12 g
Percentage yield /pə'sentɪdʒ ji:ld/	**Percentage yield** falls if some product is lost during transfer.
Empirical formula /em'pɪrɪkl 'fɔ:mjʊlə/	Ethene has the **empirical formula** CH2. CH2

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Ionic bonding /aɪˈɒnɪk ˈbɒndɪŋ/	Strong attractions between charged ions explain **ionic bonding**.
cation /ˈkætaɪən/	Sodium loses one electron to form a **cation**.
anion /ˈænaɪən/	Chlorine gains one electron to form an **anion**.
Covalent bonding /ˈkəʊvələnt ˈbɒndɪŋ/	**Covalent bonding** shares electrons between the atoms in chlorine.
Electrolysis /ɪlekˈtrɒləsɪs/	**Electrolysis** can extract aluminium from its molten compounds.
cathode /ˈkæθəʊd/	Electrons leave the hot **cathode**.
anode /ˈænəʊd/	The **anode** is the positive electrode.
oxidation /ˌɒksɪˈdeɪʃn/	**Oxidation** is the loss of electrons.
reduction /rɪˈdʌkʃn/	**Reduction** is the gain of electrons.

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