6 · Practical Skills in Chemistry II

Pearson Edexcel · International A-Level · Chemistry · Vocabulary in context

Word	Example
cation /ˈkætaɪən/	Sodium loses one electron to form a **cation**.
precipitate /prɪˈsɪpɪteɪt/	A white **precipitate** forms when silver nitrate meets chloride ions.
reagent /rɪˈeɪdʒənt/	Acidified dichromate is a **reagent** that oxidises alcohols.
titre /ˈtɪtə/	The **titre** equals the final burette reading minus the initial reading.
gradient /ˈgreɪdɪənt/	The **gradient** of a distance-time graph gives the speed.
quench /kwentʃ/	**Quench** a sample to slow its reaction before analysis.
titrate	**Titrate** the leftover iodine with sodium thiosulfate.
starch /stɑːtʃ/	Potatoes contain a lot of **starch**.
endpoint	At the **endpoint**, one more drop changes the colour for good.
half-cell /hɑːf sel/	A **half-cell** is one half of an electrochemical cell.
salt bridge /sɒlt brɪdʒ/	A **salt bridge** joins the two half-cells.
purity /ˈpjʊərɪti/	A sharp melting point shows high **purity**.
yield /jiːld/	A higher **yield** means more product is obtained.

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purification /ˌpjʊərɪfɪˈkeɪʃn/	Recrystallisation is a common **purification** method for solids.
litmus /ˈlɪtməs/	Acid turns blue **litmus** red.
ionic equation /aɪˈɒnɪk ɪˈkweɪʒn/	The **ionic equation** for neutralisation is $H+ + OH- \rightarrow H2O$. $H+ + OH- \rightarrow H2O$
clock reaction /klɒk rɪˈækʃn/	A **clock reaction** changes appearance after a measurable delay.
burette /bjʊəˈret/	A **burette** delivers measured volumes during a titration.
measuring cylinder /ˈmeʒərɪŋ ˈsɪlɪndə/	A **measuring cylinder** measures a rough volume of liquid.
water bath	Heat the test tube in a **water bath** at 80 °C. 80 °C
activation energy /ˌæktɪˈveɪʃn ˈenədʒi/	A match gives the **activation energy** to start the fire.
equilibrium /ˌiːkwɪˈlɪbrɪəm/	The book is in **equilibrium** because the forces balance.
saturated solution /ˈsætʃəreɪtɪd səˈluːʃn/	A **saturated solution** cannot dissolve any more solid.
indicator /ˈɪndɪkeɪtə/	The **indicator** changes colour at the end point.
methyl orange /ˈmiːθaɪl ˈɒrɪndʒ/	**Methyl orange** is red in acid and yellow in alkali.
phenolphthalein /ˈfenəlfθæliːn/	**Phenolphthalein** is pink in a sufficiently alkaline solution.

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concordant /kən'kɔ:dənt/	**Concordant** titres agree closely with one another.
solubility /ˌsɒlju:'bɪlɪti/	Sugar has a high **solubility** in water.
fume cupboard /fju:m 'kʌbəd/	A **fume cupboard** extracts hazardous vapours from the working space.
excess /ek'ses/	Add the metal in **excess**, so all the acid reacts.
soluble /'sɒlju:bl/	Barium hydroxide is more **soluble** than magnesium hydroxide.
impurity	An **impurity** lowers the melting point of a solid.
recrystallisation /rɪˌkrɪstəlaɪ'zeɪʃn/	**Recrystallisation** can purify a solid using hot and cold solvent.
solvent /'sɒlvənt/	Water is the **solvent** in salt water.
melting temperature /'meltɪŋ 'temprɪtʃə/	Compare the measured **melting temperature** with a reference value.
stoichiometry /ˌstəʊɪkɪ'ɒmətri/	**Stoichiometry** tells us how many moles of each substance react.
Thiele tube	A **Thiele tube** filled with oil heats the sample evenly.
capillary tube	Pack a few crystals into a **capillary tube** for the melting point test.

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