

6.2.2 Clock reactions and activation energy

Name: _____ Class: _____ Date: _____

Total: 19 marks

KEY WORDS clock reaction • activation energy • initial rate

Objective

Run the clock-reaction experiment and its arithmetic: the thiosulfate's role, why volumes are held identical, completing the reciprocal-and-log table, and the activation energy from the Arrhenius gradient. Original practice for WCH16 Q3 (Summer 2025; CP9b/CP10).

1 Key knowledge and worked examples

■ What the thiosulfate does

It reacts with iodine as it forms ($2\text{S}_2\text{O}_3^{2-} + \text{I}_2 \rightarrow \text{S}_4\text{O}_6^{2-} + 2\text{I}^-$), **delaying the colour change** until a fixed small amount of iodine has been produced; then the starch flashes **colourless** → **blue-black**. Same waiting amount every run, so $1/t \propto$ initial rate. Not: "to stop the reaction" – the scheme ignores quenching language.

■ Why a burette for every solution

Keeping the **total volume identical** keeps the reactant concentrations identical between runs – the credited reason. "More accurate" is a fallback, not the answer. Both tubes equilibrate to the water bath before mixing; record the temperature of the mixture.

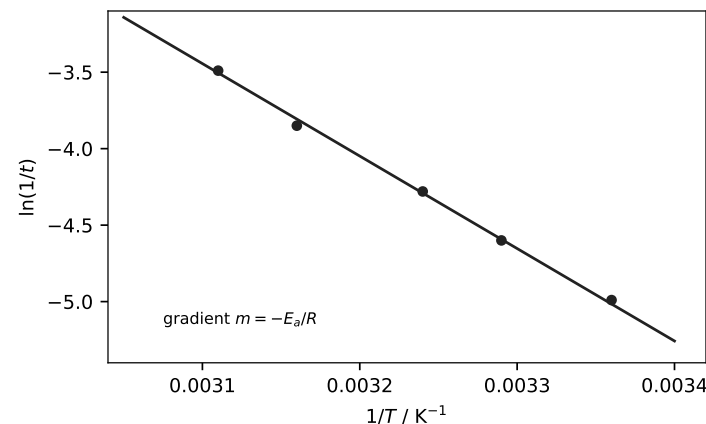
■ Table discipline

$T(\text{K}) = T(^{\circ}\text{C}) + 273$; $1/T$ to 3 s.f.; $1/t$ to 3 s.f.; $\ln(1/t)$ to 2 d.p. Check pairings row by row – a 316 K row belongs to 43 $^{\circ}\text{C}$, not 49 $^{\circ}\text{C}$.

■ The gradient pays in kJ mol^{-1}

$$\ln(1/t) = -\frac{E_a}{R} \cdot \frac{1}{T} + \text{constant}$$

Gradient $m = -E_a/R$, so $E_a = -mR = 6045 \times 8.31 = 50\,234 \text{ J mol}^{-1} = 50.2 \text{ kJ mol}^{-1}$ – positive, kJ mol^{-1} , sign carried by the gradient.



2 Practice

2.1 Explain the role of thiosulfate ions in the clock reaction, in two points the scheme credits. [2]

2.2 Why is the starch delivered from a burette rather than a measuring cylinder, even though its exact concentration does not matter? [2]

2.3 Complete: at 31 $^{\circ}\text{C}$ ($T = 304 \text{ K}$, $t = 99 \text{ s}$), calculate $1/t$ and $\ln(1/t)$. [2]

2.4 A graph of $\ln(1/t)$ against $1/T$ has gradient -7210 K . Calculate E_a in kJ mol^{-1} . [2]

3 Exam-style questions

3.1 For the peroxodisulfate–iodide clock: state the colour change at the endpoint; explain why both boiling tubes stand in the water bath until their temperatures match before mixing; and state what is plotted to give a straight line. [4]

3.2 The table gives $49\text{ }^{\circ}\text{C} \rightarrow 322\text{ K} \rightarrow 0.00311\text{ K}^{-1} \rightarrow t = 33\text{ s}$, and $25\text{ }^{\circ}\text{C} \rightarrow 298\text{ K} \rightarrow 0.00336\text{ K}^{-1} \rightarrow t = 147\text{ s}$. Complete both rows' $1/t$ and $\ln(1/t)$ values, and state the factor by which the rate changes over this range. [4]

3.3 A student's plotted gradient is -6045 K with $R = 8.31\text{ J mol}^{-1}\text{ K}^{-1}$. Calculate E_a , and explain why quoting the answer as -50.2 kJ mol^{-1} loses the mark. [3]
