4 · Chemical changes

AQA · GCSE · Chemistry · Vocabulary in context

Word	Example
positive ion /ˈpɒzɪtɪv ˈaɪɒn/	A metal atom loses electrons to form a **positive ion**.
reactivity series /rɪəkˈtɪvɪti ˈsɪəriːz/	The **reactivity series** lists metals from most to least reactive.
soluble /ˈsɒljuːbl/	Salt is **soluble** in water.
hydroxide /haɪˈdrɒksaɪd/	Sodium **hydroxide** is a strong alkali.
salt /sɒlt/	Sodium chloride is a **salt**.
displacement /dɪˈspleɪsmənt/	In a **displacement** reaction, chlorine pushes bromine out of a salt.
compound /ˈkɒmpaʊnd/	Water is a **compound** of hydrogen and oxygen.
oxide /ˈɒksaɪd/	Magnesium **oxide** forms when magnesium burns.
oxidation /ˌɒksɪˈdeɪʃn/	**Oxidation** is the loss of electrons.
reduction /rɪˈdʌkʃn/	**Reduction** is the gain of electrons.
electron /ɪˈlektrɒn/	An **electron** has a negative charge.
base /beɪs/	Sodium hydroxide is a **base**.
neutralisation /ˌnjuːtrəlaɪˈzeɪʃn/	**Neutralisation** of an acid with a base gives a salt and water.
chloride /ˈklɔːraɪd/	Sodium **chloride** is common salt.

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nitrate /ˈnaɪtreɪt/	Potassium **nitrate** is used in fertilisers.
sulfate /ˈsʌlfeɪt/	Copper **sulfate** is a blue salt.
carbonate /ˈkɑːbəneɪt/	Calcium **carbonate** fizzes with acid.
insoluble /ɪnˈsɒljuːbl/	Sand is **insoluble**: it does not dissolve in water.
filter /ˈfɪltə/	**Filter** the mixture to remove the extra solid.
filtrate /ˈfɪltreɪt/	The **filtrate** is the clear liquid that passes through the filter paper.
crystallise /ˈkrɪstəlaɪz/	Leave the solution to **crystallise** on a windowsill.
crystal	Each **crystal** of salt is a small cube.
universal indicator /ˌjuːnɪˈvɜːsl ˈɪndɪkeɪtə/	**Universal indicator** turns red in acid and purple in strong alkali.
titration /taɪˈtreɪʃn/	A **titration** finds the exact volume of acid needed.
pipette /pɪˈpet/	A **pipette** transfers an accurately measured volume.
conical flask	Swirl the **conical flask** during the titration.
burette /bjʊəˈret/	A **burette** delivers measured volumes during a titration.
meniscus /ˈmenɪskəs/	Read the volume at the bottom of the **meniscus**.
concordant titres /kənˈkɔːdənt ˈtɪtəz/	**Concordant titres** are within 0.10 cm^3 of each other. 0.10 cm^3

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strong acid /strɒŋ ˈæsɪd/	Hydrochloric acid is a **strong acid**.
ionised	A strong acid is fully **ionised** in water.
weak acid /wiːk ˈæsɪd/	Vinegar contains a **weak acid**.
ionic compound /aɪˈɒnɪk ˈkɒmpaʊnd/	Sodium chloride is an **ionic compound**.
molten /ˈməʊltn/	**Molten** salt conducts electricity because its ions can move.
electrolyte /ɪˈlektrəlaɪt/	Salt water is an **electrolyte**: ions carry the current.
lattice /ˈlætɪs/	The ions in salt sit in a regular **lattice**.
electrolysis /ɪlekˈtrɒləsɪs/	**Electrolysis** of brine makes chlorine.
cation /ˈkætaɪən/	Sodium loses one electron to form a **cation**.
cathode /ˈkæθəʊd/	Electrons leave the hot **cathode**.
anion /ˈænaɪən/	Chlorine gains one electron to form an **anion**.
anode /ˈænəʊd/	The **anode** is the positive electrode.
inert electrode /ɪˈnɜːt ɪˈlektrəʊd/	Platinum is an **inert electrode**.
cryolite /ˈkraɪəlaɪt/	**Cryolite** lowers the melting point, which saves energy.
halide /ˈhælaɪd/	Lead bromide is a **halide**: a compound of a halogen.

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halogen /ˈhælədʒn/	Chlorine is a **halogen**.
lighted splint	A **lighted splint** goes out with a pop in hydrogen.
glowing splint /ˈgləʊɪŋ splɪnt/	A **glowing splint** relights in oxygen.
litmus paper	Damp **litmus paper** turns red in acid.
half equation /hɑːf ɪˈkweɪʒn/	A **half equation** shows electrons being lost or gained.