2 · Bonding, structure and the properties of matter

AQA · GCSE · Chemistry · Vocabulary in context

Word	Example
graphene /ˈgræfiːn/	**Graphene** is a single layer of carbon atoms.
electrostatic force /ɪˌlektrəʊˈstætɪk fɔːs/	An **electrostatic force** pulls positive and negative charges together.
ionic bond /aɪˈɒnɪk bɒnd/	An **ionic bond** is the attraction between oppositely charged ions.
covalent bond /ˈkəʊvələnt bɒnd/	A **covalent bond** is a shared pair of electrons.
metallic bond /məˈtælɪk bɒnd/	A **metallic bond** attracts electrons to neighbouring positive ions.
delocalised electron /dɪˈlɒkəlaɪzd ɪˈlektrɒn/	A **delocalised electron** can move through the whole metal.
lone pair /ləʊn peə/	A **lone pair** is a pair of outer electrons that is not shared.
ionic lattice /aɪˈɒnɪk ˈlætɪs/	Salt has an **ionic lattice**: a regular pattern of positive and negative ions.
line representation /laɪn ˌreprɪˈzenteɪʃn/	In a **line representation**, each line shows one covalent bond.
intermolecular force /ˌɪntəməˈlekjʊlə fɔːs/	An **intermolecular force** acts between separate molecules.
charge carrier /tʃɑːdʒ ˈkærɪə/	In metals, electrons act as the **charge carrier**.

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alloy /ˈælɔɪ/	Brass is an **alloy** containing copper and zinc.
diamond /ˈdaɪəmənd/	**Diamond** is the hardest natural material.
giant covalent structure /ˈdʒaɪənt ˈkəʊvələnt ˈstrʌktʃə/	Diamond has a **giant covalent structure**.
graphite /ˈgræfaɪt/	Pencils contain **graphite**.
fullerene /ˈfʊlərən/	A **fullerene** is a ball of carbon atoms.
carbon nanotube /ˈkɑːbən ˌnænəʊˈtjuːb/	A **carbon nanotube** is a tiny tube made of carbon atoms.
• **nanoparticle** /ˌnænəʊˈpɑːtɪkl/	A **nanoparticle** is only a few hundred atoms across.
surface area /ˈsɜːfɪs ˈeərɪə/	Smaller solid pieces provide more **surface area** for reaction.
volume /ˈvɒljuːm/	The **volume** of the box is $2\ \text{m} \times 1\ \text{m} \times 1\ \text{m} = 2\ \text{m}^3$. $2\ \text{m} \times 1\ \text{m} \times 1\ \text{m} = 2\ \text{m}^3$

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