

9.2 The Ideal Gas Law

Name: _____ Class: _____ Date: _____

Total: 11 marks

KEY WORDS ideal gas law 理想气体定律 • temperature 温度 • pressure 压强
• ideal gas 理想气体 • heat 热量

Objective

Build the skills to answer exam questions on **the ideal gas law**.

You must be able to:

- state and apply the **ideal gas law** 理想气体定律 $PV = nRT$ (and $PV = Nk_B T$)
- analyse changes in pressure, volume, and temperature of a fixed amount of gas
- explain when a real gas behaves approximately as an **ideal gas**

1 Worked examples

Study these first. Each one shows the method for a question type used later.

■ The ideal gas law

$PV = nRT$, with $R = 8.31 \text{ J mol}^{-1} \text{ K}^{-1}$ and T in kelvin. For a **fixed** amount of gas, $\frac{PV}{T}$ is constant.

■ The combined gas law

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}.$$

At constant T , doubling P halves V (Boyle's law).

■ When a real gas is ideal

Low pressure and high temperature — the particles are far apart and their forces are weak.

2 Practice

2.1 A gas at constant temperature has its pressure doubled. State what happens to its volume. [1]

2.2 Find the volume of 2.0 mol of gas at 300 K and $1.0 \times 10^5 \text{ Pa}$ (take $R = 8.31 \text{ J mol}^{-1} \text{ K}^{-1}$). [2]

2.3 State two conditions under which a real gas behaves like an ideal gas. [2]

3 Exam-style questions

3.1 The ideal gas law is [1]

- | | |
|---|---|
| A | B |
| C | D |
- A** $PV = nRT$
B $P = \rho gh$
C $PV = \frac{1}{2}mv^2$
D $V = IR$

3.2 A fixed amount of gas is heated at constant volume. Its pressure [1]

- | | |
|---|---|
| A | B |
| C | D |
- A** falls
B rises
C stays constant
D halves

3.3 A sealed rigid 0.020 m^3 container holds gas at $2.0 \times 10^5 \text{ Pa}$ and 300 K. It is heated to 450 K.

(a) Find the new pressure. [3]

(b) State the assumption you made about the volume. [1]