

15.3 Emission and Absorption Spectra

Name: _____ Class: _____ Date: _____

Total: 10 marks

KEY WORDS emission 发射 • absorption 吸收 • absorption spectrum 吸收光谱
• emission spectrum 发射光谱 • energy levels 能级

Objective

Build the skills to answer exam questions on **emission and absorption spectra**.

You must be able to:

- explain how discrete energy levels produce line **emission** 发射 and **absorption** 吸收 spectra
- relate each spectral line to a specific electron transition
- describe how spectra identify the elements in a gas or a star

1 Worked examples

Study these first. Each one shows the method for a question type used later.

■ Line spectra

Because energy levels are **discrete**, an atom emits or absorbs only certain photon energies — giving sharp **lines**, not a continuous band.

■ Emission vs absorption

- **Emission** — bright lines from photons **released** as electrons drop down.
- **Absorption** — dark lines where photons are **absorbed** to lift electrons up.

■ A fingerprint

Each element has a unique set of lines, so spectra identify the elements in a gas or star.

2 Practice

2.1 Explain why atoms produce line spectra rather than continuous ones. [2]

2.2 State the difference between an emission and an absorption spectrum. [2]

2.3 State how a spectrum can identify an element. [1]

3 Exam-style questions

3.1 Line spectra are evidence that atomic energy levels are [1]

A	B
C	D

A continuous

B quantized

C all zero

D random

3.2 A spectrum of dark lines on a bright background is produced by [1]

A	B
C	D

A emission

B absorption

C reflection

D diffraction

3.3 The light from a star shows dark lines that match those of hydrogen.

(a) Name this type of spectrum. [1]

(b) State what it tells us about the star. [2]