5. Kinetics

Starter Quiz · 课前小测

AP Chemistry Name: _____ Score: _____

1. A reactant drops from 0.9 M to 0.6 M in 5 s. The average rate (in M/s)?

   _____ M/s

2. For rate $= k[A]^2$, doubling $[A]$ multiplies the rate by...

   ☐ 4

   ☐ 2

   ☐ 8

3. If a plot of $\ln[A]$ versus time is a straight line, the reaction is...

   ☐ first order

   ☐ second order

   ☐ zero order

4. An intermediate is a species that is...

   ☐ made in one step and used up in a later one

   ☐ present at the start and end

   ☐ never involved in the reaction

5. Select **both** conditions a collision needs to cause a reaction.

   ☐ enough energy

   ☐ correct orientation

   ☐ a colour change

   Tick every correct option.

6. Reactants at 40 kJ, peak at 100 kJ. The activation energy (in kJ)?

   _____ kJ

7. The rate-determining step is the...

   ☐ slowest elementary step

   ☐ fastest elementary step

   ☐ last step

8. If the slow first step is $NO_2 + NO_2 \rightarrow$ products, the rate law is...

   ☐ rate $= k[NO_2]^2$

   ☐ rate $= k[NO_2]$

   ☐ rate $= k$

9. An intermediate is allowed to remain in the final rate law.

   ☐ True ☐ False

10. A first-order reaction starts at 16 M with a 10 s half-life. Concentration after 20 s (in M)?

    _____ M

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Answers · 答案

AP Chemistry

1. **0.06 M/s**

   Rate $= (0.9 - 0.6)/5 = 0.06$ M/s.

2. **4**

   Second order means $2^2 = 4$ times the rate.

3. **first order**

   A linear $\ln[A]$-versus-time plot indicates first order.

4. **made in one step and used up in a later one**

   An intermediate forms then is consumed, so it cancels overall.

5. **enough energy; correct orientation**

   A successful collision needs both sufficient energy and the right orientation.

6. **60 kJ**

   $E_a = 100 - 40 = 60$ kJ (peak minus reactants).

7. **slowest elementary step**

   The slowest step is the bottleneck controlling the overall rate.

8. **rate** $= k[\mathbf{NO_2}]^2$

   Two NO_2 react in the slow step, so it is second order.

9. **False**

   It must be substituted out using the fast equilibrium.

10. **4 M**

    Two half-lives give $16 \to 8 \to 4$ M.