

3 Forces between molecules and states of matter – Part 1 —Answers

Answers

Work through these after you have tried the questions yourself.

2.1 weaker (intermolecular forces are much weaker than covalent or ionic bonds).

2.2 London dispersion forces, then dipole–dipole forces, then hydrogen bonding.

2.3 water molecules attract each other by strong hydrogen bonds, which take much more energy to break; methane has only weak dispersion forces.

2.4 London dispersion forces.

2.5 the forces **between** the molecules are broken; the bonds inside each water molecule stay intact.