

9.3 Gibbs Free Energy and Thermodynamic Favorability

Name: _____ Class: _____ Date: _____ **Total: 11 marks****KEY WORDS** thermodynamically favorable 热力学有利 • gibbs free energy 吉布斯自由能

Objective

Build the skills to answer exam questions on **Gibbs free energy** and thermodynamic favorability.**You must be able to:**

- use $\Delta G = \Delta H - T\Delta S$
- interpret the sign of ΔG (favorable if < 0)
- find the temperature where a reaction becomes favorable

1 Worked examples

Study these first. Each one shows the method for a question type used later — follow the steps and you can do the Practice and Exam-style questions yourself.

■ The Gibbs equation

$$\Delta G = \Delta H - T\Delta S,$$

with T in **kelvin**. A reaction is **thermodynamically favorable** (spontaneous) when $\Delta G < 0$.

■ The four cases

- $\Delta H < 0$, $\Delta S > 0$: **always** favorable ($\Delta G < 0$ at all T).
- $\Delta H > 0$, $\Delta S < 0$: **never** favorable.
- $\Delta H < 0$, $\Delta S < 0$: favorable at **low** T .
- $\Delta H > 0$, $\Delta S > 0$: favorable at **high** T .

■ A worked ΔG

 $\Delta H = -100 \text{ kJ}$, $\Delta S = +50 \text{ J/K} = 0.050 \text{ kJ/K}$, $T = 300 \text{ K}$:

$$\Delta G = -100 - 300(0.050) = -100 - 15 = -115 \text{ kJ (favorable)}.$$

■ Crossover temperature

The reaction switches at $\Delta G = 0$: $T = \frac{\Delta H}{\Delta S}$. Watch **units** — convert ΔS to kJ/K to match ΔH .

2 Practice

Now apply the methods above.

2.1 Write the Gibbs free energy equation. [1]**2.2** A reaction has $\Delta G = -40 \text{ kJ}$. Is it favorable? [1]**2.3** For $\Delta H = -50 \text{ kJ}$, $\Delta S = +0.100 \text{ kJ/K}$, $T = 300 \text{ K}$, find ΔG . [2]

3 Exam-style questions

3.1 A reaction is thermodynamically favorable when ΔG is [1]

A	B
C	D

- A positive
B negative
C zero
D equal to ΔH

3.2 A reaction has $\Delta H = +30 \text{ kJ}$ and $\Delta S = +0.10 \text{ kJ/K}$.

(a) Is it favorable at low or high temperature? [1]

(b) Find the temperature at which $\Delta G = 0$. [3]

3.3 A reaction has $\Delta H < 0$ and $\Delta S < 0$. Explain at what temperatures it is favorable and why. [2]