

7.3 The Reaction Quotient and Equilibrium Constant

Name: _____ Class: _____ Date: _____

Total: 10 marks

KEY WORDS reaction quotient 反应商

Objective

Build the skills to answer exam questions on the **reaction quotient** Q and its comparison to K .

You must be able to:

- calculate Q with the same form as K but current (non-equilibrium) concentrations
- compare Q to K to predict the direction of shift
- state what $Q = K$ means

1 Worked examples

Study these first. Each one shows the method for a question type used later — follow the steps and you can do the Practice and Exam-style questions yourself.

■ The reaction quotient

Q has the **same expression** as K but uses the **current** concentrations (not necessarily at equilibrium).

■ Comparing Q and K

- $Q < K$: too few products $\rightarrow$ shifts **forward** (toward products);
- $Q > K$: too many products $\rightarrow$ shifts **reverse** (toward reactants);
- $Q = K$: **at equilibrium**, no shift.

■ A worked comparison

If $K = 10$ and the current $Q = 2$, then $Q < K$, so the reaction proceeds **forward** to make more product until Q rises to 10.

■ Why it works

Q measures "how far along" the mixture is. The reaction always moves to bring Q toward K .

2 Practice

Now apply the methods above.

2.1 What does $Q = K$ mean? [1]

2.2 If $Q < K$, which way does the reaction shift? [1]

2.3 How does the expression for Q differ from that for K ? [1]

3 Exam-style questions

3.1 If $Q > K$, the reaction shifts [1]

A	B
C	D

A forward (toward products)

B reverse (toward reactants)

C it is at equilibrium

D it stops

3.2 For a reaction, $K = 50$. A mixture is analyzed and $Q = 120$.

(a) Compare Q and K . [1]

(b) State the direction of shift, with a reason. [2]

3.3 For $A \rightleftharpoons B$ with $K = 4$, a mixture has $[A] = 2$ M and $[B] = 2$ M. Find Q and predict the shift. [3]