

6.5 Energy of Phase Changes

Name: _____ Class: _____ Date: _____

Total: 9 marks

KEY WORDS heat of fusion 熔化热 • phase change 相变 • heat of vaporization 汽化热
• heat 热量

Objective

Build the skills to answer exam questions on the **energy of phase changes**.

You must be able to:

- use the **heat of fusion** 熔化热 and **heat of vaporization** 汽化热
- explain the flat plateaus on a heating curve
- combine heating and phase-change steps

1 Worked examples

Study these first. Each one shows the method for a question type used later — follow the steps and you can do the Practice and Exam-style questions yourself.

■ Phase-change energy

Melting and boiling require energy to overcome intermolecular forces, at **constant temperature**:

$$q = n \Delta H_{\text{fus}} \quad (\text{melting}), \quad q = n \Delta H_{\text{vap}} \quad (\text{boiling}).$$

■ Why the temperature is flat

During a phase change the added energy goes into **separating particles**, not raising their kinetic energy — so the temperature holds steady (the plateau on a heating curve).

■ Combining steps

To take ice below 0°C to steam, add up: warm ice ($mc\Delta T$) + melt ($n\Delta H_{\text{fus}}$) + warm water ($mc\Delta T$) + boil ($n\Delta H_{\text{vap}}$) + warm steam.

■ A worked plateau

Boiling always needs **more** energy than melting for the same substance ($\Delta H_{\text{vap}} > \Delta H_{\text{fus}}$), because boiling fully separates the particles.

2 Practice

Now apply the methods above.

2.1 During melting, what happens to the temperature? [1]

2.2 Write the equation for the heat to melt n moles of a solid. [1]

2.3 Which needs more energy per mole: melting or boiling (same substance)? [1]

3 Exam-style questions

3.1 During a phase change, added energy goes into [1]

A	B
C	D

A raising the temperature

B overcoming intermolecular forces

C breaking covalent bonds

D speeding up the particles

3.2 2.0 mol of ice melts. $\Delta H_{\text{fus}} = 6.0 \text{ kJ/mol}$.

(a) Find the heat needed. [2]

(b) State what happens to the temperature during melting. [1]

3.3 Explain why the boiling plateau on a heating curve is longer than the melting plateau for the same substance and heating rate. [2]