

Solutions

Each answer shows the steps, then the **result**.

2.1

- A proton (H^+) donor

2.2

- $\text{KNO}_3 + \text{H}_2\text{O}$ (salt + water)

2.3

- F^-

3.1 B

- a proton acceptor

3.2

(a)

- H_2O is the acid (donates H^+); NH_3 is the base (accepts)

(b)

- $\text{NH}_3/\text{NH}_4^+$ and $\text{H}_2\text{O}/\text{OH}^-$

3.3

- $\text{H}_2\text{SO}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + 2\text{H}_2\text{O}$