

2.5 Lewis Diagrams

Name: _____ Class: _____ Date: _____

Total: 10 marks

KEY WORDS lone pairs 孤对电子 • valence electrons 价电子 • lewis diagram 路易斯结构
• octet 八隅体

Objective

Build the skills to answer exam questions on **Lewis diagrams** — drawing the shared and lone pairs of a molecule.

You must be able to:

- count total **valence electrons** 价电子 for a molecule or ion
- arrange bonding pairs and **lone pairs** 孤对电子 to satisfy the octet
- use double or triple bonds when needed

1 Worked examples

Study these first. Each one shows the method for a question type used later — follow the steps and you can do the Practice and Exam-style questions yourself.

■ Counting valence electrons

Add the valence electrons of every atom; for an **ion**, add one electron per negative charge and subtract one per positive charge. For CO_2 : $4 + 2(6) = 16$.

■ Building the skeleton

Put the least electronegative atom in the centre (usually not H), connect atoms with single bonds, then add lone pairs to complete octets (H needs only 2).

■ Using multiple bonds

If the central atom is short of an octet, form **double or triple bonds**. For CO_2 , each O shares a double bond with C: $\text{O} = \text{C} = \text{O}$, all octets satisfied, 16 electrons used.

■ Checking

Count electrons: each bond is 2, each lone pair is 2. The total must equal your valence-electron count, and every atom (except H) should have an octet.

2 Practice

Now apply the methods above.

2.1 Count the total valence electrons in H_2O . [1]

2.2 Count the total valence electrons in NH_3 . [1]

2.3 How many lone pairs are on the oxygen in a water molecule? [1]

3 Exam-style questions

3.1 The number of valence electrons to use in the Lewis diagram of NO_3^- is [1]

A	B
C	D

A 22

B 23

C 24

D 5

3.2 Draw the Lewis diagram of CO_2 .

(a) State the total number of valence electrons. [1]

(b) Describe the bonding (single/double/triple) between C and each O. [2]

3.3 For N_2 , state the total valence electrons and the bond order (single/double/triple) needed to satisfy both octets. [3]