

2.3 Structure of Ionic Solids

Name: _____ Class: _____ Date: _____

Total: 11 marks**KEY WORDS** lattice 晶格 • lattice energy 晶格能 • ionic solid 离子固体

Objective

Build the skills to answer exam questions on the **structure of ionic solids**.**You must be able to:**

- describe an ionic solid as a **lattice** 晶格 of alternating + and – ions
- explain properties (high melting point, brittle, conducts only when molten/dissolved) from the structure
- link **lattice energy** 晶格能 to ion charge and size

1 Worked examples

Study these first. Each one shows the method for a question type used later — follow the steps and you can do the Practice and Exam-style questions yourself.

■ The ionic lattice

An ionic solid is a giant regular array (**lattice**) of cations and anions, each surrounded by ions of opposite charge. Strong electrostatic forces act in all directions.

■ Properties from structure

- High melting point** — many strong ionic attractions must be broken.
- Brittle** — a shifted layer puts like charges together, and they repel, cracking the crystal.
- Conducts electricity only when molten or dissolved** — ions are then free to move; in the solid they are locked in place.

■ Lattice energy

The **lattice energy** (energy to separate the ions) rises with **larger ion charges** and **smaller ion sizes** ($E \propto \frac{q_1q_2}{r}$). MgO has a much larger lattice energy than NaCl.

■ Predicting melting point

More lattice energy $\Rightarrow$ higher melting point. Compare compounds by their ion charges first, then their sizes.

2 Practice

Now apply the methods above.

2.1 Why do ionic solids have high melting points? [1]**2.2** Why does solid NaCl not conduct electricity, but molten NaCl does? [2]**2.3** State two factors that increase lattice energy. [2]

3 Exam-style questions

3.1 An ionic solid conducts electricity when [1]

A	B
C	D

- A** solid
B molten or dissolved in water
C cooled
D never

3.2 Compare NaF and MgO.

(a) State which has the larger lattice energy, with a reason. [2]

(b) Predict which has the higher melting point. [1]

3.3 Explain why ionic solids are brittle, in terms of what happens when a layer of ions is shifted. [2]