

1.3 Elemental Composition of Pure Substances

Name: _____ Class: _____ Date: _____

Total: 14 marks

KEY WORDS empirical formula 实验式 • molecular formula 分子式
 • percent composition 百分组成 • molar mass 摩尔质量 • moles 摩尔

Objective

Build the skills to answer exam questions on **the elemental composition of pure substances**.

You must be able to:

- calculate the **percent composition** 百分组成 by mass of each element
- find an **empirical formula** 实验式 from mass percentages
- find a **molecular formula** 分子式 from the empirical formula and molar mass
- convert masses to moles and divide by the smallest

1 Worked examples

Study these first. Each one shows the method for a question type used later — follow the steps and you can do the Practice and Exam-style questions yourself.

■ Percent composition

$$\% \text{ element} = \frac{\text{mass of that element}}{\text{total mass}} \times 100\%.$$

In water ($M = 18$), oxygen is $16/18 \times 100 \approx 89\%$.

■ Empirical formula

Convert each element's mass to moles, then divide by the smallest. For 40% C, 6.7% H, 53.3% O (per 100 g): C = 3.33, H = 6.7, O = 3.33 mol; dividing by 3.33 gives CH₂O.

■ Molecular formula

Divide the true molar mass by the empirical mass for the multiple n . Glucose: empirical CH₂O (30), molar mass 180, so $n = 6$, giving C₆H₁₂O₆.

2 Practice

Now apply the methods above.

2.1 In water (H₂O, $M = 18$), what percent of the mass is oxygen (16)? [2]

2.2 State what an empirical formula gives. [1]

2.3 A compound has empirical mass 30 and molar mass 90. Find the multiple n . [2]

2.4 After dividing by the smallest you get a ratio 1 : 1.5. State the whole-number ratio.[1]

3 Exam-style questions

3.1 The empirical formula of a compound gives [1]

A	B
C	D

A the exact number of atoms in a molecule

B the simplest whole-number ratio of atoms

C the molar mass

D the percent composition

3.2 A compound is 75% carbon and 25% hydrogen by mass (C = 12, H = 1). Its empirical formula is [1]

A	B
C	D

A CH

B CH₂

C CH₄

D C₂H

3.3 A compound is 40% sulfur and 60% oxygen by mass ($S = 32$, $O = 16$).

(a) Find the empirical formula. [3]

(b) State the method you used. [1]

3.4 A compound has empirical formula CH_2 (empirical mass 14) and a molar mass of 56 g/mol.

(a) Find the multiple n . [1]

(b) Write the molecular formula. [1]