

1.1 Moles and Molar Mass

Name: _____ Class: _____ Date: _____

Total: 16 marks**KEY WORDS** mole 摩尔 • molar mass 摩尔质量 • avogadro's number 阿伏伽德罗常数

Objective

Build the skills to answer exam questions on **moles and molar mass**.**You must be able to:**

- use the **mole** 摩尔 as a fixed count, 6.022×10^{23} particles (**Avogadro's number** 阿伏伽德罗常数)
- use the **molar mass** 摩尔质量 (g/mol) to convert between grams and moles
- convert with $n = \frac{m}{M}$, and moles to particles by multiplying by Avogadro's number
- find the molar mass of a compound by adding atomic masses

1 Worked examples

Study these first. Each one shows the method for a question type used later — follow the steps and you can do the Practice and Exam-style questions yourself.

■ Moles from mass

$$n = \frac{m}{M}$$

36 g of water ($M = 18$ g/mol) is $n = 36/18 = 2$ mol.

■ Molar mass of a compound

Add the atomic masses. For H_2O : $2(1) + 16 = 18$ g/mol. For CO_2 : $12 + 2(16) = 44$ g/mol.

■ Moles to particles

Multiply by Avogadro's number: 2 mol is $2 \times 6.022 \times 10^{23} \approx 1.2 \times 10^{24}$ molecules.

2 Practice

Now apply the methods above.

2.1 How many moles are in 44 g of CO_2 ($M = 44$ g/mol)? [2]

2.2 Find the molar mass of water, H_2O (H = 1, O = 16). [2]

2.3 Find the mass of 3 mol of carbon atoms ($M = 12$ g/mol). [2]

2.4 State the number of particles in one mole. [1]

3 Exam-style questions

3.1 One mole of any substance contains [1]

- | | |
|---|---|
| A | B |
| C | D |
- A** 1000 particles
B 6.022×10^{23} particles
C 6.022×10^{-23} particles
D 12 particles

3.2 The molar mass of CO_2 (C = 12, O = 16) is [1]

- | | |
|---|---|
| A | B |
| C | D |
- A** 28 g/mol
B 44 g/mol
C 12 g/mol
D 60 g/mol

3.3 A sample contains 88 g of CO_2 ($M = 44 \text{ g/mol}$).

(a) Find the number of moles. [2]

(b) Find the number of molecules (use 6.022×10^{23}). [2]

3.4 A student needs 0.5 mol of NaCl ($M = 58.5 \text{ g/mol}$).

(a) Find the mass required. [2]

(b) State the units of molar mass. [1]