

7.11 Introduction to Solubility Equilibria

Name: _____ Class: _____ Date: _____ **Total: 12 marks****KEY WORDS** molar solubility 摩尔溶解度 • solubility product 溶度积

Objective

Build the skills to answer exam questions on **solubility equilibria** and K_{sp} .**You must be able to:**

- write the **solubility product** 溶度积 K_{sp} expression
- relate **molar solubility** 摩尔溶解度 to K_{sp}
- calculate one from the other

1 Worked examples

Study these first. Each one shows the method for a question type used later — follow the steps and you can do the Practice and Exam-style questions yourself.

■ The K_{sp} expression

For a sparingly soluble salt dissolving, e.g. $\text{AgCl}(s) \rightleftharpoons \text{Ag}^+ + \text{Cl}^-$:

$$K_{sp} = [\text{Ag}^+][\text{Cl}^-].$$

The solid is omitted.

■ Molar solubility

Let s be the moles of salt that dissolve per litre. Each formula unit gives ions according to its formula: for AgCl , $[\text{Ag}^+] = [\text{Cl}^-] = s$, so $K_{sp} = s^2$.

■ A worked calculation

If $K_{sp}(\text{AgCl}) = 1.8 \times 10^{-10}$, then $s = \sqrt{K_{sp}} = \sqrt{1.8 \times 10^{-10}} = 1.3 \times 10^{-5} \text{ M}$.

■ Salts giving more ions

For $\text{CaF}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{F}^-$: $[\text{Ca}^{2+}] = s$, $[\text{F}^-] = 2s$, so $K_{sp} = s(2s)^2 = 4s^3$.

2 Practice

Now apply the methods above.

2.1 Write the K_{sp} expression for $\text{BaSO}_4 \rightleftharpoons \text{Ba}^{2+} + \text{SO}_4^{2-}$. [1]**2.2** For a 1:1 salt with molar solubility s , write K_{sp} in terms of s . [1]**2.3** If $K_{sp} = 4 \times 10^{-10}$ for a 1:1 salt, find s . [2]

3 Exam-style questions

3.1 For $\text{AgCl} \rightleftharpoons \text{Ag}^+ + \text{Cl}^-$ with molar solubility s , $K_{sp} =$ [1]

- **A** s
- **B** s^2
- **C** $2s$
- **D** $4s^3$

3.2 $\text{PbCl}_2 \rightleftharpoons \text{Pb}^{2+} + 2\text{Cl}^-$ has molar solubility s .(a) Write $[\text{Pb}^{2+}]$ and $[\text{Cl}^-]$ in terms of s . [2](b) Write K_{sp} in terms of s . [2]

3.3 $K_{sp}(\text{CaF}_2) = 3.2 \times 10^{-11}$, with $K_{sp} = 4s^3$. Find the molar solubility s . [3]

Solutions

2.1 $K_{sp} = [\text{Ba}^{2+}][\text{SO}_4^{2-}]$.

2.2 $K_{sp} = s^2$.

2.3 $s = \sqrt{4 \times 10^{-10}} = 2 \times 10^{-5} \text{ M}$.

3.1 B — $K_{sp} = s^2$ for a 1:1 salt.

3.2 (a) $[\text{Pb}^{2+}] = s$, $[\text{Cl}^-] = 2s$. (b) $K_{sp} = s(2s)^2 = 4s^3$.

3.3 $4s^3 = 3.2 \times 10^{-11}$; $s^3 = 8 \times 10^{-12}$; $s = 2 \times 10^{-4} \text{ M}$.